



# Craft and Activity Book

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# Halloween Fun Beginners Activity Books

**Tony J. Tallarico**



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 $2\text{NaCl} + \text{F}_2 \rightarrow 2\text{NaF} + \text{Cl}_2$   
 $4\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$   
 $5\text{Pb} + 2\text{HCl} \rightarrow 2\text{H}_2 + \text{PbCl}_2$   
 $2\text{Al} + 3\text{K}_2\text{SO}_4 \rightarrow 4\text{K} + 3\text{Al}_2\text{O}_3 + 6\text{S}$   
 $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$   
 $8\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$

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web oct 2 2019 let's start by balancing the oxygen atoms to do this make the oxygen atoms as six on either side of the chemical equation  $2\text{Fe}_2\text{O}_3 \rightarrow 4\text{Fe} + 3\text{CO}_2$  on towards the next step now step 3 once you are done with this check out if the number of atoms for the other elements is also equal on both the sides

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 $1\text{ N}_2 + \text{F}_2 \rightarrow 2\text{NF}_3$   
 $2\text{C}_6\text{H}_{10}\text{O}_2 + \text{CO}_2 + \text{H}_2\text{O} \rightarrow 3\text{HBr} + \text{KHCO}_3 + \text{H}_2\text{O} + \text{KBr} + \text{CO}_2 + \text{GaBr} + \text{NaBr} + \text{SnO}_4 + \text{NaSO}_3 + \text{Ga}_2\text{SO}_3 + 3\text{NF}_3 + 3\text{SnF}_2 + \text{N}_2\text{O}_3$  using the equation from problem 2 above answer the following

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 $1\text{ N}_2 + 3\text{F}_2 \rightarrow 2\text{NF}_3$   
 $2\text{C}_6\text{H}_{10}\text{O}_2 + 17\text{O}_2 \rightarrow 12\text{CO}_2 + 10\text{H}_2\text{O}$   
 $3\text{HBr} + \text{KHCO}_3 + \text{H}_2\text{O} + \text{KBr} + \text{CO}_2 + 2\text{GaBr}_3 + 3\text{Na}_2\text{SO}_3 + \text{Ga}_2\text{SO}_3 + 6\text{NaBr} + 3\text{SnO}_2 + 3\text{NF}_3 + 3\text{SnF}_2 + \text{N}_2\text{O}_3$  using the following equation  $2\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{Na}_2\text{SO}_4$

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web may 28 2020 write the balanced equation for the reaction and calculate how much  $\text{sio}_2$  is required to produce 3.00 kg of  $\text{sic}$  answer  $\text{ce sio}_2 \text{ 3c} \rightarrow \text{sic 2co}$

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