

2 homework point(s)

Galvanic Cells

Using tabulated standard reduction potentials from your text, calculate the standard cell potential, E°_{cell} (always positive for a galvanic cell), based on the following (unbalanced) reaction:



$$E^\circ_{\text{cell}} = 0.036 \text{ V}$$

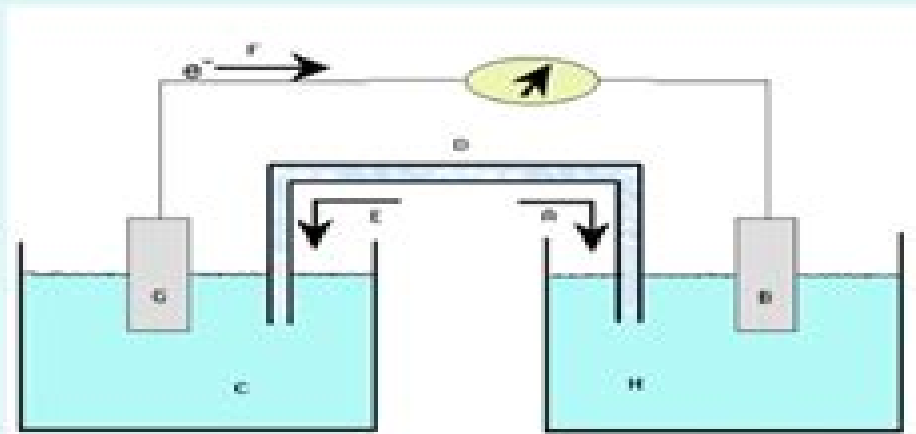
of Table A5.5, p. A24 Zumdahl "Chemical Principles" 7th ed.

1 pt

You are correct. Computer's answer now shown above.
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A galvanic cell based on the above reaction is constructed according to the generic sketch of a galvanic cell shown below:



Choose the label which best describes the following:

C: Solution containing Cr^{3+} , H^+ and $\text{Cr}_2\text{O}_7^{2-}$

B: $\text{Cl}_2(\text{g})/\text{Pt}$ cathode

The overall, balanced, spontaneous reaction occurring in this (standard) cell is:



False: The salt bridge disrupts the electrical circuit while permitting mixing of the solutions in the anode and cathode compartments.

False: The salt bridge is labeled "D".

True: The solution in tube "D" must contain an electrolyte for the cell to operate.

True: The salt bridge is intended to allow counter ions to flow between the two cell compartments.

False: Tube "D" allows Cr^{3+} ions to migrate readily to the other compartment and to react directly with the chlorine gas.

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