



Long March 3b User Manual

Joost J.L.M. Bierens



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web jan 29 2020 the ionic form of this reaction is 2 na aq 2 cl aq cu 2 aq so 4 2 aq 2 na aq so 4 2 aq cucl 2 s the sodium ions and sulfate ion are the spectator ions in this reaction they appear unchanged in both the product and reactant side of the equation

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web the six no 3 aq ions and the six na aq ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation 3ba 2 aq 2po 4 3 aq rightarrow ba 3 po 4 2 s

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web these spectator ions ions whose presence is required to maintain charge neutrality are neither chemically nor physically changed by the process and so they may be eliminated from the equation to yield a more succinct representation called a

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oxidation number or composition during a reaction products

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water h_2O h^+aq oh^-aq h_2o l the spectator ions in koh aq and $\text{hno}_3 \text{aq}$ are k^+aq and $\text{no}_3^- \text{aq}$ how to identify and list the

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web the ag^+aq and cl^-aq ions become agcl s but the na^+aq ions and the $\text{no}_3^- \text{aq}$ ions stay as na^+aq ions and $\text{no}_3^- \text{aq}$ ions these two ions are examples of spectator ions ions that do nothing in the overall course of a chemical reaction

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web aug 8 2022 a spectator ion is an ion that does not take part in the chemical reaction and is found in solution both before and after the reaction the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction

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web the new format requires students to write balanced chemical equations showing only the reacting substances excluding spectator ions for three different sets of reactants and to answer a short question requiring no calculator about each reaction the questions are intended to test the students understanding of the meaning of the reactions

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web jun 2 2023 the spectator ions are cut out from a reaction hence the net ionic equation for the above reaction would be $2\text{Cl}^- \text{aq} + \text{Cu}^{2+} \text{aq} \rightarrow \text{Cu}^{2+} \text{aq} + 2\text{Cl}^- \text{aq}$ these spectator ions are not actively involved in the reaction the existence of spectator ions affects the Debye length

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