



CHEMISTRY

CHEAT SHEET

EQUILIBRIUM

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Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time having reached Equilibrium. Equilibrium reactions are represented by the symbol \rightleftharpoons .

<p>What is Equilibrium?</p> <p>Equilibrium is a state of balance between two or more opposing processes.</p> <p>In chemistry, it refers to a state where the concentrations of reactants and products remain constant over time.</p> <p>Equilibrium is reached when the rate of the forward reaction is equal to the rate of the reverse reaction.</p> <p>Equilibrium is a dynamic process, meaning that the reactions are still occurring, but the concentrations of the species remain constant.</p>	<p>Equilibrium Constant</p> <p>The equilibrium constant (K) is a measure of the extent to which a reaction proceeds. It is defined as the ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	<p>Graph showing concentration vs. time for a reaction at equilibrium.</p> <p>Reactant concentration: $[A]_0 - [A]_t = [A]_{eq}$</p> <p>Product concentration: $[C]_0 + [C]_t = [C]_{eq}$</p>
<p>Equilibrium Constant, K_c</p> <p>The equilibrium constant K_c is calculated using the concentrations of the species in the reaction mixture.</p> <p>For the reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant K_c is given by:</p> $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	<p>Equilibrium Constant, K_p</p> <p>The equilibrium constant K_p is calculated using the partial pressures of the gaseous species in the reaction mixture.</p> <p>For the reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant K_p is given by:</p> $K_p = \frac{(P_C)^c (P_D)^d}{(P_A)^a (P_B)^b}$	<p>Equilibrium Constant, K</p> <p>The equilibrium constant K is a general term that can refer to either K_c or K_p, depending on the context of the reaction.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant K is given by:</p> $K = \frac{[C]^c [D]^d}{[A]^a [B]^b}$
<p>Le Chatelier's Principle: the effect of changing conditions on equilibrium</p>	<p>Equilibrium Constant, K</p> <p>The equilibrium constant K is a general term that can refer to either K_c or K_p, depending on the context of the reaction.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant K is given by:</p> $K = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	<p>Equilibrium Constant, K</p> <p>The equilibrium constant K is a general term that can refer to either K_c or K_p, depending on the context of the reaction.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant K is given by:</p> $K = \frac{[C]^c [D]^d}{[A]^a [B]^b}$
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Guide Chemical Equilibrium

Michael Hoelscher



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