

Chemistry Worksheet on Chapter 3 Electrochemistry -Set 4

Q-1: Consider the electrochemical cell

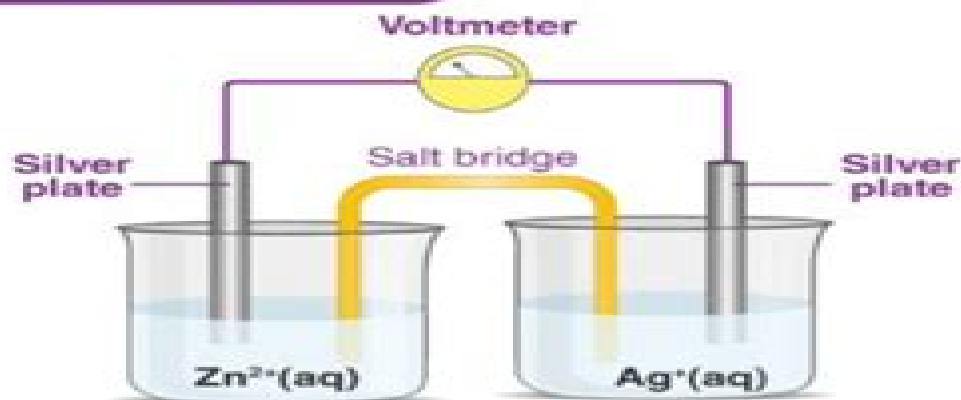


Where 'M' is a metal. At 298 K, the standard reduction potentials are $E^\circ(M^{2+}(aq)/M(s)) = -0.12\text{ V}$, $E^\circ(M_2(s)/M(s)) = -0.36\text{ V}$ and the temperature coefficient is $1.5 \times 10^{-4}\text{ V K}^{-1}$. At this temperature the standard enthalpy change for the overall cell reaction, $\Delta_r H^\circ$ is _____ kJ/mol. (Round off to two decimal places) (Faraday constant $F = 96500\text{ C/mol}$)

Q-2: Define the term "electrochemical cell." What happens if the applied external potential exceeds the E°_{cell} of an electrochemical cell?

Q-3: Consider the diagram and respond to the following questions.

ELECTROCHEMICAL CELL



- Which is cathode and which is anode?
- What will happen if the salt bridge is removed?
- What is the direction of electron flow?

Q-4: At infinite dilution, the equivalent conductance of CH_3COONa , HCl and CH_3COOH are 91, 426 and 391 mho cm^2 respectively at 25°C . The eq. conductance of $NaCl$ at infinite dilution will be (in $\text{S cm}^2 \text{eq}^{-1}$)

Holt Electrochemistry Review Answers

Robert R. Redfield



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