

ACIDS, BASES AND SALTS

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| Definition of an acid An acid is a compound which when dissolved in water produces hydrogen ions as the only positively charged ions | What causes the acidic properties of acids? The hydrogen ions (H^+) cause the acidic properties and these are formed in the presence of water. |
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| Another term to refer to an acid An acid is called a <u>proton donor</u> | Why an acid is also called a proton donor? It's because an acid provides protons or hydrogen ions (H^+) to other substances during the reaction. |
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| Substance to which an acid provides protons BASE | A base is then called PROTON ACCEPTOR | Why the base is called a proton acceptor? Because the base accepts hydrogen ions from acids | Equation for the reaction between acid and base $H^+_{(aq)} + OH^-_{(aq)} \longrightarrow H_2O_{(l)}$ |
|--|---|---|---|

| Common laboratory acids | | These three common laboratory acids are also called | Why these acids are also called mineral acids? |
|-------------------------|-----------|---|---|
| Hydrochloric acid | HCl | MINERAL ACIDS | They are derived from mineral salts i.e chlorides for HCl, sulphates for H_2SO_4 and nitrates for HNO_3 |
| Sulphuric acid | H_2SO_4 | | |
| Nitric acid | HNO_3 | | |

| Other mineral acids known | | Mineral salts from which the acid is derived | Organic acids known | Naturally occurring acids |
|---------------------------|-----------|--|------------------------------|---|
| Sulphurous acid | H_2SO_3 | Derived from sulphites | Ethanoic acid (CH_3COOH) | CITRIC ACID from lemons |
| Carbonic acid | H_2CO_3 | Derived from carbonates | | TARTARIC ACID from grapes |
| Phosphoric acid | H_3PO_4 | Derived from phosphates | Methanoic acid ($HCOOH$) | ACETIC ACID from vinegar |
| Nitrous acid | HNO_2 | Derived from nitrites | | LACTIC ACID from sour milk |
| | | | | Hydrochloric acid from digestive juices |

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| Whenever an acid is dissolved in water, it produces HYDROGEN IONS | Term given to the number of hydrogen ions produced by one molecule of an acid BASICITY OF AN ACID | Definition of basicity of an acid BASICITY of an acid is the <u>number of hydrogen ions</u> produced by <u>one molecule</u> of an acid <u>in aqueous solution</u> . |
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Guide Acids Bases And Salts

Marco Cascella



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