

CH 4. RADICALS

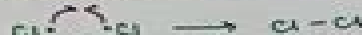
1. INITIATION - FORM RADICALS



2. PROPAGATION - RADICAL REACTANT → RADICAL PRODUCT



3. TERMINATION - 2 RADICALS → NO RADICALS



$$\Delta H_{\text{rxn}} = \sum (\text{bonds broken}) - \sum (\text{bonds formed}) - \text{form bond}$$



F_2 - most reactive
 Cl_2 - reactive, not selective
 Br_2 - selective, picky
 I_2 - not reactive, endo.



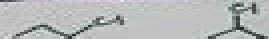
RESONANCE + SUBST. ↑ RAD. STABILITY
 $\text{Allyl} \cdot > \text{R}-\text{C}(\text{H})-\text{R} > \text{R}-\text{C}(\text{H})_2-\text{R} > \text{R}-\text{C}(\text{H})_3$
 $3^\circ > 2^\circ > 1^\circ$

CH 5. STEREOCHEMISTRY

CHIRAL - Not symmetrical



CONST. ISOMER - connect. diff.



ACHIRAL - SYMMETRY



ENANTIOMER - Invert chiral centers



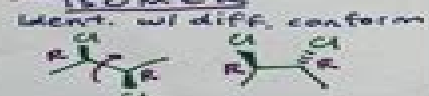
DIASTEREOMER - Invert some, not all, chiral centers



MESO - ACHIRAL W/ CHIRAL CENTERS



CONFORMATIONAL ISOMER



CH 6. NUCLEOPHILIC SUBSTITUTION

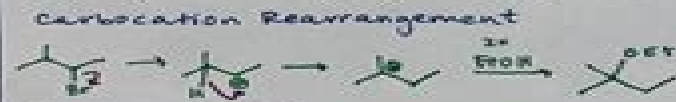


Nuc.
 Lone pairs
 (b) charge
 (b) or δ^+ charge

LEAVING GROUP: $\text{I} > \text{Br} > \text{Cl} \gg \text{F}$ or OH

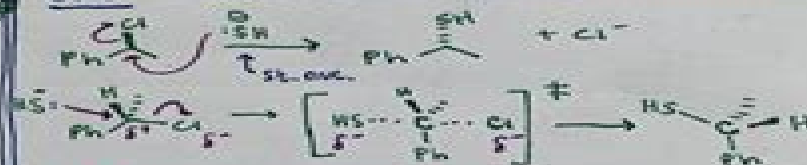


SN1 → Racemic



* Weak Nucleophile, Rate = $k[\text{subst.}]$

SN2 → Invert chiral center



* Strong Nucleophile
 1° or methyl

$$\text{rate} = k[\text{substrate}][\text{nucleophile}]$$

Guide Quick Notes Reactions Iitjee Organic Chemistry

Richard Charles Whitfield



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