

# Thermochemistry Worksheet Answers

## Useful Values:

Heat of Fusion for Water:  $\Delta H_{\text{fus}} = 6 \text{ kJ/mol}$

**Common error: not converting to moles**

Heat of Vaporization for Water:  $\Delta H_{\text{vap}} = 45 \text{ kJ/mol}$

**Common error: same**

1 calorie = 4.18 Joules

Material	Specific Heat in Joules/g/K	in Cal/g/K
Water (ice)	2.05	0.490
Water (liquid)	4.18	1.00
Water (gas)	2.08	0.498

**Common error here:  
using moles instead of grams.**

1. How much energy does it take to warm the following substances the listed amount?

- |   |                                   |
|---|-----------------------------------|
| a. 20g of ice from $-45^{\circ}\text{C}$ to $-15^{\circ}\text{C}$ ?         | <b>294 calories / 1230 Joules</b> |
| b. 20g of water from $25^{\circ}\text{C}$ to $55^{\circ}\text{C}$ ?         | <b>600 calories / 2508 Joules</b> |
| c. 20g of water vapor from $130^{\circ}\text{C}$ to $160^{\circ}\text{C}$ ? | <b>299 calories / 1249 Joules</b> |

2. To melt ice, one must bring ice to the freezing point ( $0^{\circ}\text{C}$ ), and then thaw it into water. How much heat would it take to melt...

- |  |                 |                   |
|--|-----------------|-------------------|
| a. 30g of ice at $0^{\circ}\text{C}$ ?   | <b>10 kJ</b>    | <b>(10,000 J)</b> |
| b. 30g of ice at $-40^{\circ}\text{C}$ ? | <b>12.46 kJ</b> | <b>(12,460 J)</b> |

3. To boil water, one must raise it to the boiling point ( $100^{\circ}\text{C}$ ), and then boil it into steam. How much heat would it take to boil...

- |  |                 |                         |
|--|-----------------|-------------------------|
| a. 40g of water at $100^{\circ}\text{C}$ ? | <b>100 kJ</b>   | <b>(100,000 Joules)</b> |
| b. 40g of water at $90^{\circ}\text{C}$ ?  | <b>101.7 kJ</b> | <b>(101,672 Joules)</b> |
| c. 40g of water at $0^{\circ}\text{C}$ ?   | <b>116.7 kJ</b> | <b>(116,720 Joules)</b> |
| d. 40g of ice at $0^{\circ}\text{C}$ ?     | <b>130.1 kJ</b> | <b>(130,053 Joules)</b> |
| e. 40g of ice at $-80^{\circ}\text{C}$ ?   | <b>136.6 kJ</b> | <b>(136,613 Joules)</b> |

4. If 700 calories of heat raise 400g of Phonium by  $20^{\circ}\text{C}$ , what is the specific heat of Phonium? **0.0875 cal/g/K**

5. How much heat energy is stored in 2g of ice at  $-30^{\circ}\text{C}$ ? (remember: use Kelvin) **238 cal / 996 J**

6. How much heat energy is stored in 2g of Phonium at  $70^{\circ}\text{C}$ ? (Assume Phonium is a solid at  $70^{\circ}\text{C}$ ) **1672 cal**

7. If 100g of water at  $10^{\circ}\text{C}$  comes in contact with \_\_\_\_\_, what will the final temperature be?

(Remember, you have the specific heat of Phonium from question 4.)

- |  |   |
|--|---|
| a. 100g of water at $90^{\circ}\text{C}$   | <b><math>50^{\circ}\text{C}</math></b>    |
| b. 100g of Phonium at $90^{\circ}\text{C}$ (Hint: the water and Phonium end at the same temperature) | <b><math>16.44^{\circ}\text{C}</math></b> |
| <b>Common error: 6.44 degrees is the temperature change, and the water gets warmer.</b>              |   |
| c. 20g of Phonium at $90^{\circ}\text{C}$  | <b><math>11.38^{\circ}\text{C}</math></b> |
| d. 1g of Phonium at $90^{\circ}\text{C}$   | <b><math>10.07^{\circ}\text{C}</math></b> |

8. If 450g of Phakelite at  $100^{\circ}\text{C}$  is added to 1000g of water at  $20^{\circ}\text{C}$ , the temperature of the water increases by  $5^{\circ}$ . What is the specific heat of Phakelite? **0.148 cal/g/K**

8. If 450g of Nonsensium at  $500^{\circ}\text{C}$  is added to 1000g of water at  $20^{\circ}\text{C}$ , the temperature of the water increases by  $25^{\circ}$ . What is the specific heat of Nonsensium? **0.122 cal/g/K**

# Guided And Study Workbook Thermochemistry Answers

**M Woodhall**



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