

General Chemistry II Jasperse
Chemical equilibria. Extra Practice Problems

General Types/Groups of problems:

Equilibrium Conceptual	p1	Using Iee: Generic, Then Real But Simple Numbers	p8
Writing the Equilibrium Constant	p3	Solving for K given Initial and at Least one Equilibrium Concentration	p9
Manipulations of K: Reversing or Multiplying	p5	Solving for Equilibrium Concentrations Using Iee, Given K	p10
Solving for K Given All Equilibrium Concentrations	p6	LeChatelier's Principle, Q, and Responses to Disrupted Equilibria	p12
Solving for an Equilibrium Concentration Given K and Other Equilibrium Concentrations	p7	Answers	p15

Big-Picture Introductory Conceptual Questions

- Which of the following is true for a chemical reaction at equilibrium?
 - only the forward reaction stops
 - only the reverse reaction stops
 - both the forward and reverse reactions stop
 - the rate constants for the forward and reverse reactions are equal
 - the rates of the forward and reverse reactions are equal
- Which of the following is true regarding the **concentration of products**, for a chemical reaction **that is already at equilibrium**, assuming no disruptions to the equilibrium?
 - The concentrations of products will not change because there are no more reactants.
 - The concentrations of products will not change because the limiting reagent is gone.
 - The concentrations of products will not change because the forward and reverse rates are equal.
 - The concentrations of products will change continually because of reversibility.
- Which of the following are equal for a chemical system at equilibrium? If all are equal, answer E.
 - the **concentrations** of reactant and products are equal
 - the **rate constants** for the forward and reverse reactions are equal
 - the **time** that a particular atom or molecule spends as a reactant and product are equal
 - the **rate** of the forward and reverse reaction
 - all** of the above are equal
- A chemical equilibrium may be established by starting a reaction with _____.
 - reactants only.
 - products only.
 - equal quantities of reactants and products.
 - any quantities of reactants and products.
 - all the above
- An equilibrium that strongly favors products has _____.
 - a value of $K \ll 1$.
 - a value of $K \gg 1$.
 - a value of $Q \gg 1$.
 - a value of $Q \ll 1$.
 - $K = Q$.

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Arshad Iqbal



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Table of Contents Guided Practice Problems Chemistry Ch 16

1. Understanding the eBook Guided Practice Problems Chemistry Ch 16
 - The Rise of Digital Reading Guided Practice Problems Chemistry Ch 16
 - Advantages of eBooks Over Traditional Books
2. Identifying Guided Practice Problems Chemistry Ch 16
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Guided Practice Problems Chemistry Ch 16
 - User-Friendly Interface
4. Exploring eBook Recommendations from Guided Practice Problems Chemistry Ch 16
 - Personalized Recommendations
 - Guided Practice Problems Chemistry Ch 16 User Reviews and Ratings
 - Guided Practice Problems Chemistry Ch 16 and Bestseller Lists
5. Accessing Guided Practice Problems Chemistry Ch 16 Free and Paid eBooks
 - Guided Practice Problems Chemistry Ch 16 Public Domain eBooks
 - Guided Practice Problems Chemistry Ch 16 eBook Subscription Services
 - Guided Practice Problems Chemistry Ch 16 Budget-Friendly Options
6. Navigating Guided Practice Problems Chemistry Ch 16 eBook Formats
 - ePub, PDF, MOBI, and More
 - Guided Practice Problems Chemistry Ch 16 Compatibility with Devices
 - Guided Practice Problems Chemistry Ch 16 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Guided Practice Problems Chemistry Ch 16
 - Highlighting and Note-Taking Guided Practice Problems Chemistry Ch 16
 - Interactive Elements Guided Practice Problems Chemistry Ch 16

8. Staying Engaged with Guided Practice Problems Chemistry Ch 16
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Guided Practice Problems Chemistry Ch 16
9. Balancing eBooks and Physical Books Guided Practice Problems Chemistry Ch 16
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Guided Practice Problems Chemistry Ch 16
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Guided Practice Problems Chemistry Ch 16
 - Setting Reading Goals Guided Practice Problems Chemistry Ch 16
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Guided Practice Problems Chemistry Ch 16
 - Fact-Checking eBook Content of Guided Practice Problems Chemistry Ch 16
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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