

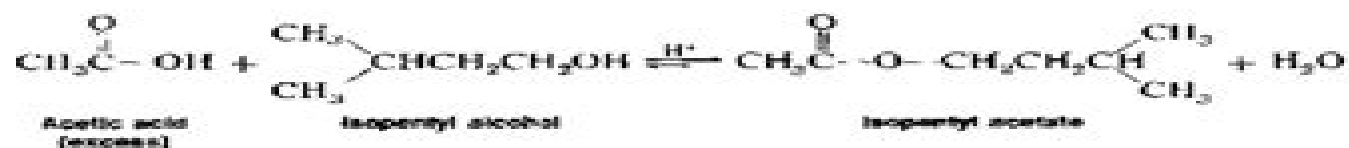
Experiment #14:

ISOPENTYL ACETATE

Data Table for Exp #14

Name	Mass (g)
Mass of empty 5mL conical vial	25.021 g
Mass of 5mL conical vial + Isopentyl Alcohol	25.834 g
Mass of Isopentyl Alcohol	0.813 g
Mass of Isopentyl Acetate (Theoretical)	1.20 g
Mass of 5mL conical vial + Isopentyl Acetate (collected)	25.943 g
Mass of Isopentyl Acetate (Experiment)	0.922 g
% Yield of Isopentyl Acetate	76.83%
Boiling Point of Isopentyl Acetate	140°C

Calculation:

Theoretical Yield of Isopentyl Acetate

$$1 \text{ ml of Isopentyl Alcohol} \times \frac{0.813 \text{ g of Isopentyl Alcohol}}{1 \text{ ml Isopentyl Alcohol}} \times \frac{1 \text{ mol of Isopentyl Alcohol}}{88.2 \text{ g of Isopentyl Alcohol}} = 9.22 \times 10^{-3} \text{ mol of Isopentyl Alcohol}$$

$$9.22 \times 10^{-3} \text{ mol of Isopentyl Alcohol} \times \frac{1 \text{ mol of Isopentyl Acetate}}{1 \text{ mol Isopentyl Alcohol}} \times \frac{130.19 \text{ g Isopentyl Acetate}}{1 \text{ mol of Isopentyl Acetate}} = 1.20 \text{ g Isopentyl Acetate (Theoretical)}$$

$$\% \text{Yield of Isopentyl Acetate} = \frac{\text{experiment result}}{\text{theoretical result}} = \frac{0.922 \text{ g}}{1.20 \text{ g}} = 76.83\%$$

Isopentyl Acetate Lab Report

Judd E. Hollander



Isopentyl Acetate Lab Report:

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