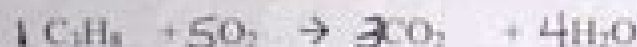


Limiting Reagent Worksheet #1

1. Given the following reaction: (Balance the equation first!)



a) If you start with 14.8 g of C_3H_8 and 3.44 g of O_2 , how many grams of H_2O are produced?

$$\frac{14.8 \text{ g C}_3\text{H}_8}{44.097 \text{ g C}_3\text{H}_8} \times \frac{1 \text{ mol C}_3\text{H}_8}{1 \text{ mol C}_3\text{H}_8} \times \frac{4 \text{ mol H}_2\text{O}}{1 \text{ mol C}_3\text{H}_8} \times \frac{18.015 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = \frac{1066.488}{44.097} = 24.19 \text{ g}$$

$$\frac{3.44 \text{ g O}_2}{31.998 \text{ g O}_2} \times \frac{1 \text{ mol O}_2}{1 \text{ mol O}_2} \times \frac{4 \text{ mol H}_2\text{O}}{5 \text{ mol O}_2} \times \frac{18.015 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = \frac{247.8564}{159.99} = 1.549 \text{ g}$$

b) Which is the limiting reagent (which reactant created less H_2O)? _____

How many grams of the non-limiting reagent are actually used? _____

How many grams of the non-limiting reagent are in excess (left over)? _____

Following equation:



Limiting Reactants Study Guide Answer Key

**ONS, June Eilers, Martha
Langhorne, Regina Fink**



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Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst. Carbon nanotubes are a recently discovered allotrope of carbon. Three other allotropes of carbon, buckyballs, graphite, and diamond, are illustrated at the left, as is the molecule methane CH_4 from which nanotubes and buckyballs can be made. The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet. The study of chemistry also follows from the simple to the more complex, and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement. **Class 8-12**

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