

7 • Liquids & Solids

Released FRQs

1979 D

Butane, chloroethane, acetone, and 1-propanol all have approximately the same molecular weights. Data on their boiling points and solubilities in water are listed in the table below.

Compound	Formula	Boiling Point (°C)	Solubility in water
Butane	$\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3$	0	insoluble
Chloroethane	$\text{CH}_3\text{CH}_2\text{Cl}$	12	insoluble
Acetone	$\begin{array}{c} \text{O} \\ \\ \text{CH}_3\text{CCH}_3 \end{array}$	56	completely miscible
1-Propanol	$\text{CH}_3\text{CH}_2\text{CH}_2\text{OH}$	97	completely miscible

On the basis of dipole moments (molecular polarities) and/or hydrogen bonding, explain in a qualitative way the differences in the

- boiling points of butane and chloroethane.
- water solubilities of chloroethane and acetone.
- water solubilities of butane and 1-propanol.
- boiling points of acetone and 1-propanol.

Answer 1:

- Butane is nonpolar; chloroethane is polar. Intermolecular forces of attraction in liquid chloroethane are larger due to dipole-dipole attraction; thus a higher boiling point for chloroethane.
- Both chloroethane and acetone are polar. However, acetone forms hydrogen bonds to water much more effectively than chloroethane does, resulting in greater solubility of acetone in water.
- Butane is non-polar and cannot form hydrogen bonds; 1-propanol is polar and can form hydrogen bonds. 1-propanol can interact with water by both dipole-dipole forces and hydrogen bonds. Butane can interact with water by neither means. Thus, 1-propanol is much more soluble.
- Acetone molecules are attracted to each other by van der Waals attraction and dipole-dipole attraction. 1-propanol molecules show these two types of attraction. However, 1-propanol can also undergo hydrogen bonding. This distinguishing feature results in the higher boiling point of 1-propanol.

1977 D (Mandatory)

The state of aggregation of solids can be described as belonging to the following four types:

- (1) ionic
- (2) metallic
- (3) covalent network
- (4) molecular

For each of these types of solids, indicate the kinds of particles that occupy the lattice points and identify forces among these particles. How could each type of solid be identified in the laboratory?

Answer 1:

	Particles	Binding Forces	Experimental Identification
ionic	+ & - ions	electrostatic attraction	conductivity of fused salt
metallic	+ ions	electrostatic attraction between ions and electrons	conductivity of the solid

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Michael Hoelscher



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