

Chemistry 100: Final Exam Questions

Question 1: Which compound is a tertiary alcohol?
 a. 1-propanol
 b. 2-propanol
 c. 1-butanol
 d. 2-butanol
 e. 2-methyl-2-butanol

Question 2: Which of the following is the correct Lewis structure for water?
 a. 
 b. 
 c. 
 d. 

Question 3: Which of the following represents the correct Lewis structure for ammonia?
 a. 
 b. 
 c. 
 d. 

Question 4: A gas sample weighs 0.450 g. It has a volume of 100 mL at 25°C and 1 atm. The gas is composed of 80% carbon and 20% hydrogen. What is the molecular formula of the gas?
 a. C_2H_4
 b. C_2H_6
 c. C_3H_8
 d. C_4H_{10}
 e. C_5H_{12}

Question 5: The compound with the empirical formula $\text{C}_2\text{H}_4\text{O}$ has a molar mass of 88 g/mol. What is the molecular formula of the compound?
 a. $\text{C}_2\text{H}_4\text{O}$
 b. $\text{C}_4\text{H}_8\text{O}_2$
 c. $\text{C}_6\text{H}_{12}\text{O}_3$
 d. $\text{C}_8\text{H}_{16}\text{O}_4$
 e. $\text{C}_{10}\text{H}_{20}\text{O}_5$

Question 6: A sample of gas occupies 100 mL at 25°C and 1 atm. It is cooled to 0°C and 1 atm. What is the new volume of the gas?
 a. 100 mL
 b. 125 mL
 c. 150 mL
 d. 175 mL
 e. 200 mL

Question 7: Which of the following represents the correct Lewis structure for carbon dioxide?
 a. 
 b. 
 c. 
 d. 

Question 8: The density of a gas is 1.25 g/L at 25°C and 1 atm. What is the molecular formula of the gas?
 a. C_2H_4
 b. C_2H_6
 c. C_3H_8
 d. C_4H_{10}
 e. C_5H_{12}

Question 9: The density of a gas is 1.25 g/L at 25°C and 1 atm. What is the molecular formula of the gas?
 a. C_2H_4
 b. C_2H_6
 c. C_3H_8
 d. C_4H_{10}
 e. C_5H_{12}

Question 10: A gas sample weighs 0.450 g. It has a volume of 100 mL at 25°C and 1 atm. The gas is composed of 80% carbon and 20% hydrogen. What is the molecular formula of the gas?
 a. C_2H_4
 b. C_2H_6
 c. C_3H_8
 d. C_4H_{10}
 e. C_5H_{12}

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Norman C. Gysbers, Patricia Henderson

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