

THE UNIVERSITY OF CHICAGO
DEPARTMENT OF CHEMISTRY
RESEARCH REPORT

RESEARCH REPORT

1. **Introduction**
The purpose of this report is to describe the results of the experiments conducted during the summer of 1964. The experiments were designed to investigate the effect of temperature on the rate of reaction between hydrogen peroxide and potassium permanganate in the presence of ceric sulfate as a catalyst. The reaction was studied at various temperatures ranging from 10°C to 40°C. The rate of reaction was determined by measuring the volume of oxygen gas evolved over a period of time. The results of the experiments are presented in the following sections.

2. **Experimental Procedure**
The reaction was carried out in a 250 ml Erlenmeyer flask. A solution of potassium permanganate (KMnO₄) was prepared by dissolving a known weight of the salt in distilled water. A solution of ceric sulfate (Ce(SO₄)₂) was also prepared. A solution of hydrogen peroxide (H₂O₂) was prepared by diluting a stock solution of 30% H₂O₂ with distilled water. The reaction mixture was prepared by combining the solutions of KMnO₄, Ce(SO₄)₂, and H₂O₂ in the Erlenmeyer flask. The flask was then placed in a water bath at the desired temperature. The volume of oxygen gas evolved was measured by the displacement of water in an inverted graduated cylinder.

3. **Results and Discussion**
The results of the experiments are shown in the following table. The rate of reaction was determined by measuring the volume of oxygen gas evolved over a period of time. The rate of reaction was found to increase with increasing temperature. The rate of reaction was also found to be dependent on the concentration of the reactants. The rate of reaction was found to be independent of the concentration of the catalyst, ceric sulfate. The results of the experiments are consistent with the proposed mechanism for the reaction.

4. **Conclusion**
The results of the experiments show that the rate of reaction between hydrogen peroxide and potassium permanganate in the presence of ceric sulfate as a catalyst increases with increasing temperature. The rate of reaction is also dependent on the concentration of the reactants. The rate of reaction is independent of the concentration of the catalyst, ceric sulfate. The results of the experiments are consistent with the proposed mechanism for the reaction.

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