

**its second half equivalence point?**

- A.  $[\text{CO}_3^{2-}] = \text{maximum value}$
- B.  $[\text{HCO}_3^-] = \text{maximum value}$
- C.  $[\text{H}_2\text{CO}_3] = [\text{HCO}_3^-]$
- D.  $[\text{HCO}_3^-] = [\text{CO}_3^{2-}]$
- E.  $[\text{H}_2\text{CO}_3] = [\text{CO}_3^{2-}]$

Answer: D.  $[\text{HCO}_3^-] = [\text{CO}_3^{2-}]$

**4. The  $K_a$  for HCN is  $6.2 \times 10^{-10}$ , If there is a solution of 2M HCN, what concentration of NaCN is needed in order for the pH to be 9.2?**

- A. 3M
- B. No NaCN needs to be added
- C. 2M
- D. 1M

Answer: C. 2M

**5. NaOH is added to a 500ml of 2M acetic acid. If the  $pK_a$  value of acetic acid is approximately 4.8, what volume of 2M NaOH must be added so that the pH of the solution is 4.8?**

- A. 2L
- B. 500ml
- C. 250mL
- D. 1L

Answer: C. 250ml

**6. Consider the titration curve, in which an acid is treated with aqueous 0.5M NaOH. What is the approximate  $pK_a$  of the starting acid?**

- A. 4.2
- B. 3.0
- C. 8.9
- D. 5.0
- E. 7.0

# Gre Chemistry Study Guide

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### **Table of Contents Gre Chemistry Study Guide**

1. Understanding the eBook Gre Chemistry Study Guide
  - The Rise of Digital Reading Gre Chemistry Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Gre Chemistry Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Gre Chemistry Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Gre Chemistry Study Guide
  - Personalized Recommendations

- Gre Chemistry Study Guide User Reviews and Ratings
- Gre Chemistry Study Guide and Bestseller Lists
- 5. Accessing Gre Chemistry Study Guide Free and Paid eBooks
  - Gre Chemistry Study Guide Public Domain eBooks
  - Gre Chemistry Study Guide eBook Subscription Services
  - Gre Chemistry Study Guide Budget-Friendly Options
- 6. Navigating Gre Chemistry Study Guide eBook Formats
  - ePub, PDF, MOBI, and More
  - Gre Chemistry Study Guide Compatibility with Devices
  - Gre Chemistry Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Gre Chemistry Study Guide
  - Highlighting and Note-Taking Gre Chemistry Study Guide
  - Interactive Elements Gre Chemistry Study Guide
- 8. Staying Engaged with Gre Chemistry Study Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Gre Chemistry Study Guide
- 9. Balancing eBooks and Physical Books Gre Chemistry Study Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Gre Chemistry Study Guide
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Gre Chemistry Study Guide
  - Setting Reading Goals Gre Chemistry Study Guide
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Gre Chemistry Study Guide
  - Fact-Checking eBook Content of Gre Chemistry Study Guide

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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