

## Procedure

You will use the simulation on the following website to virtually measure and record all voltages for this experiment.

<http://web.mst.edu/~gbert/Electro/Electrochem.html>

Once you have gone to the website above you should see the interface below.



- You will be able to adjust the electrodes and solutions using the drop down menus on the right and left. The simulator will only allow you to use a solution of the metal salt corresponding to the metal electrode placed in the beaker. (For example: Copper(II) nitrate must be used with the copper electrode.)
- You can also adjust the concentration of each of the metal salt solutions in the beakers by typing the value in the box on either side of the simulator.
- To measure the voltage of any cell that you have constructed, simply click **Measure Cell Voltage** on the bottom left of the simulator. If there is a problem with the cell, the simulator will identify any errors in your construction.
- You will not be using the **New Problem** link in the or "Level" drop down menu in the lower right for this assignment.

## Part A

1. You will be measuring the cell potentials for all the possible combinations of each of the following electrodes:  $\text{Ag}/\text{Ag}^+$ ,  $\text{Cu}/\text{Cu}^{2+}$ ,  $\text{Ni}/\text{Ni}^{2+}$ ,  $\text{Pb}/\text{Pb}^{2+}$ , and  $\text{Zn}/\text{Zn}^{2+}$ . Construct the cells so that the metal salt solution in each beaker has a concentration of 0.10 M and a positive voltage is read on the voltmeter when the cell potential is measured.
2. Place a silver electrode in the beaker on the left using the top drop down menu and then select silver nitrate in the second drop down menu and type in a concentration of 0.10 into the box at the bottom.

# Lab Manual For Chem 1212l

**N Noddings**



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