

Isotopes Practice Set

- What does atomic number tell you about an element?
- What does the number used to isotopes signify?
- How can you tell isotopes apart?

For each of the following isotopes, write the number of protons, neutrons, and electrons.

	Chlorine-35	Chlorine-37
# of protons	17	17
# of neutrons	18	20
# of electrons	17	17

	Nitrogen-14	Nitrogen-15
# of protons	7	7
# of neutrons	7	8
# of electrons	7	7

	Carbon-12	Carbon-13
# of protons	6	6
# of neutrons	6	7
# of electrons	6	6

	Oxygen-16	Oxygen-18
# of protons	8	8
# of electrons	8	8
# of neutrons	8	10

Fill in the isotopes names including atomic number, chemical symbol, and mass number from the values that your partner gave and the information provided.

	Isotope: 17	Isotope: 20
# of protons	17	17
# of neutrons	18	20
# of electrons	17	17

	Isotope: 14	Isotope: 15
# of protons	7	7
# of neutrons	7	8
# of electrons	7	7

	Isotope: 12	Isotope: 13
# of protons	6	6
# of neutrons	6	7
# of electrons	6	6

	Isotope: 16	Isotope: 18
# of protons	8	8
# of neutrons	8	10
# of electrons	8	8

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 - The Rise of Digital Reading Isotope Practice Set Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Isotope Practice Set Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Isotope Practice Set Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Isotope Practice Set Answer Key
 - Personalized Recommendations
 - Isotope Practice Set Answer Key User Reviews and Ratings
 - Isotope Practice Set Answer Key and Bestseller Lists

5. Accessing Isotope Practice Set Answer Key Free and Paid eBooks
 - Isotope Practice Set Answer Key Public Domain eBooks
 - Isotope Practice Set Answer Key eBook Subscription Services
 - Isotope Practice Set Answer Key Budget-Friendly Options
6. Navigating Isotope Practice Set Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Isotope Practice Set Answer Key Compatibility with Devices
 - Isotope Practice Set Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Isotope Practice Set Answer Key
 - Highlighting and Note-Taking Isotope Practice Set Answer Key
 - Interactive Elements Isotope Practice Set Answer Key
8. Staying Engaged with Isotope Practice Set Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Isotope Practice Set Answer Key
9. Balancing eBooks and Physical Books Isotope Practice Set Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Isotope Practice Set Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Isotope Practice Set Answer Key
 - Setting Reading Goals Isotope Practice Set Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Isotope Practice Set Answer Key
 - Fact-Checking eBook Content of Isotope Practice Set Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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