

Ionic Salts Simulation Lab**Learning Goals:**

- Understand how an ionic salt dissolves describing a saturated solution microscopically and macroscopically.
- Calculate solubility in grams/100ml
- Determine why some salts are more soluble than others
- Design and carry out a virtual lab to test a hypothesis

Directions: Go to

http://phet.colorado.edu/simulations/sims.php?sim=Salts_and_Solubility and run the Salts and Solubility simulation. Take a few minutes to get use to how the simulation works. Then answer the questions below.

Questions:

- 1) Describe what you observe when you shake table salt into water. Test numbers of particles with different volumes.
 - a. When the salt dissolves, which element becomes the anion (the negatively charged atom) and which one becomes the cation (the positively charged atom)?
 - b. What is the relationship between amount of particles and amount of water? (direct or indirect).
 - c. At 5.00×10^{-23} what is the point of saturation (point where you can't dissolve any more salt)?
- 2) Fill in the chart below to see if the ratio of particles to water is constant at the saturation point.

Particles	Water Volume	Ratio (particles/volume)

- a. Calculate the constant below (show your work)

- 3) Test a variety of slightly soluble salts, noting the relationship between numbers of each ion pair. Fill in the chart below to help you see a relationship.
 - a. How many particles of each element need to be present before the salt stops dissolving (point of saturation)?

Lab 30 Solubility Of A Salt Answers

Thomas Griffiths



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