

15. (Pre-lecture) Study the following reaction:
- a. Write out the **balanced equation**  
b. Calculate  $\Delta H_{\text{rxn}}$  **kJ/mol**
16. (Pre-lecture) This reaction is thermodynamically favored at 298 K. Is it also kinetically favored?  
a. Explain why or why not.  
b. Draw a reaction coordinate diagram for this reaction. Label the axes and the transition state.
17. (Pre-lecture) Calculate the equilibrium constant  $K_{\text{eq}}$  for the reaction of  $\text{H}_2\text{O}$  and  $\text{CO}_2$  at 298 K.  
a. Write the **balanced equation**  
b. Calculate  $\Delta G_{\text{rxn}}$  **kJ/mol**  
c. Calculate  $K_{\text{eq}}$
18. (Pre-lecture) Consider the following reaction. Write the **balanced equation** and calculate  $\Delta G_{\text{rxn}}$  **kJ/mol** at 298 K.

Reactants	Products
$\text{H}_2\text{O}$	$\text{H}_2\text{O}$
$\text{CO}_2$	$\text{CO}_2$
$\text{H}_2$	$\text{H}_2$
$\text{O}_2$	$\text{O}_2$

19. (Pre-lecture) The following is a balanced equation for the reaction of  $\text{H}_2\text{O}$  and  $\text{CO}_2$  at 298 K. Calculate  $\Delta G_{\text{rxn}}$  **kJ/mol** at 298 K. Is the reaction spontaneous?  
a. Write the **balanced equation**  
b. Calculate  $\Delta G_{\text{rxn}}$  **kJ/mol**

$$\Delta G_{\text{rxn}} = \Delta H_{\text{rxn}} - T\Delta S_{\text{rxn}} = \Delta H_{\text{rxn}} - (298 \text{ K})\Delta S_{\text{rxn}}$$

20. (Pre-lecture) Consider the following reaction. Write the **balanced equation** and calculate  $\Delta G_{\text{rxn}}$  **kJ/mol** at 298 K.

$$\Delta G_{\text{rxn}} = \Delta H_{\text{rxn}} - T\Delta S_{\text{rxn}} = \Delta H_{\text{rxn}} - (298 \text{ K})\Delta S_{\text{rxn}}$$

21. (Pre-lecture) Calculate the **equilibrium constant** for the reaction of  $\text{H}_2\text{O}$  and  $\text{CO}_2$  at 298 K.  
a. Write the **balanced equation**  
b. Calculate  $\Delta G_{\text{rxn}}$  **kJ/mol**  
c. Calculate  $K_{\text{eq}}$

# Lab Conductivity Quizz With Answers

**Albert A Gayle**



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