

Ideal Gas Law

$$PV = nRT$$

↑ ↑ ↑ ↑
L mol K

atm

K

$$R = 0.08206$$

$$\frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}}$$

$$1 \text{ L} = 1000 \text{ mL}$$

$$T_K = T_C + 273$$

$$1 \text{ atm} = 760 \text{ torr} = 101.3 \text{ kPa}$$

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N Noddings



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