

Ideal Gas Law Practice Worksheet

Solve the following problems using the ideal gas law:

- 1) How many moles of gas does it take to occupy 120 liters at a pressure of 2.3 atmospheres and a temperature of 340 K?
 $PV = nRT$
 $V = 120\text{ L}$
 $P = 2.3\text{ atm}$
 $T = 340\text{ K}$
 $n = ?$
 $(2.3\text{ atm})(120\text{ L}) = n(0.0821 \frac{\text{L}\cdot\text{atm}}{\text{K}\cdot\text{mol}})(340\text{ K})$
 $n = 9.9\text{ mol}$
- 2) If I have a 50 liter container that holds 45 moles of gas at a temperature of 200° C, what is the pressure inside the container?
 $V = 50\text{ L}$
 $n = 45\text{ mol}$
 $T = 200^\circ\text{C} = 473\text{ K}$
 $P = ?$
 $(P)(50\text{ L}) = (45\text{ mol})(0.0821 \frac{\text{L}\cdot\text{atm}}{\text{K}\cdot\text{mol}})(473\text{ K})$
 $P = 30\text{ atm}$ or 4000 kPa
- 3) It is not safe to put aerosol canisters in a campfire, because the pressure inside the canisters gets very high and they can explode. If I have a 1.0 liter canister that holds 2 moles of gas, and the campfire temperature is 1400° C, what is the pressure inside the canister?
 $V = 1.0\text{ L}$
 $n = 2\text{ mol}$
 $T = 1400^\circ\text{C} = 1673\text{ K}$
 $P = ?$
 $(P)(1.0\text{ L}) = (2\text{ mol})(0.0821 \frac{\text{L}\cdot\text{atm}}{\text{K}\cdot\text{mol}})(1673\text{ K})$
 $P = 275\text{ atm} = 280\text{ atm}$
or $28,000\text{ kPa}$
- 4) How many moles of gas are in a 30 liter scuba canister if the temperature of the canister is 300 K and the pressure is 200 atmospheres?
 $V = 30\text{ L}$
 $T = 300\text{ K}$
 $P = 200\text{ atm}$
 $n = ?$
 $(200\text{ atm})(30\text{ L}) = n(0.0821 \frac{\text{L}\cdot\text{atm}}{\text{K}\cdot\text{mol}})(300\text{ K})$
 $n = 243.6\text{ mol} = 240\text{ mol}$
- 5) I have a balloon that can hold 100 liters of air. If I blow up this balloon with 3.0 moles of oxygen gas at a pressure of 1 atmosphere, what is the temperature of the balloon?
 $V = 100\text{ L}$
 $n = 3\text{ mol}$
 $P = 1\text{ atm}$
 $T = ?$
 $(1\text{ atm})(100\text{ L}) = (3\text{ mol})(0.0821 \frac{\text{L}\cdot\text{atm}}{\text{K}\cdot\text{mol}})(T)$
 $T = 406\text{ K} = 133^\circ\text{C} = 130^\circ\text{C}$

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Patrick Hurley

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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