#### GAS LAW WORKSHEET 1 KEY

 A sample of oxygen gas occupies a volume of 436.2 mL at 1.02 atm. If the temperature is held constant, what would the pressure of this gas be when the gas is compressed to 231.6 mL?

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2} \qquad \text{since } T_1 = T_2 \qquad P_1 V_1 = P_2 V_2 \qquad P_2 = P_1 \frac{V_1}{V_2} = 1.02 \text{ atm} \left( \frac{436.2 \text{ mL}}{231.6 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm} \left( \frac{100 \text{ mL}}{100 \text{ mL}} \right) = 1.92 \text{ atm$$

To check with reasoning: decreased V leads to increased P, so the volume ratio should be greater than one to yield the expected increased pressure.

If a gas originally occupying 6.75 L at 19.21 °C and 762.5 torr is compressed to give a pressure of 1.26 atm at 26.35 °C, what would the new volume be?

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2} \qquad V_2 = V_1 \left(\frac{T_2}{T_1}\right) \left(\frac{P_1}{P_2}\right) = 6.75 \text{ L} \left(\frac{299.50 \text{ K}}{292.36 \text{ K}}\right) \left(\frac{762.5 \text{ torr}}{1.26 \text{ atm}}\right) \left(\frac{1 \text{ atm}}{760 \text{ torr}}\right) = 5.51 \text{ L}$$

To check with reasoning: increased T leads to increased V, so the temperature ratio should be greater than one to yield the expected increase in volume. Increased P leads to decreased V, so the pressure ratio should be less than one to yield the expected decreased volume.

Calculate the number of grams of hydrogen sulfide gas, H<sub>2</sub>S, in a 3.6 L container at 32.6 °C and 712 mmHg.

$$PV = \frac{g}{M}RT \quad g = \frac{PVM}{RT} = \frac{712 \text{ mmHg (3.6 L) } 34.080 \frac{g}{mol}}{0.082058 \frac{L \text{ atm}}{K \text{ mol}}} \left(\frac{1 \text{ atm}}{760 \text{ mmHg}}\right) = 4.6 \text{ g H}_2S$$

4. Calculate the density of 0.625 g of carbon dioxide at 26.32 °C and 1.03 atm.

$$PV = \frac{g}{M}RT - \frac{g}{V} = \frac{PM}{RT} = \frac{1.03 \text{ atm} \left(44.010 \frac{g}{\text{mol}}\right)}{0.082058 \frac{L \text{ atm}}{K \text{ mol}} \left(299.47 \text{ K}\right)} = 1.84 \text{ g/L CO}_2$$

# **Ideal Gas Laws Ws Answers**

**Bernard Liengme, Keith Hekman** 

#### **Ideal Gas Laws Ws Answers:**

Survey of Science History & Concepts Parent Lesson Plan ,2013-08-01 Survey of Science History Concepts Course Description Students will study four areas of science Scientific Mathematics Physics Biology and Chemistry Students will gain an appreciation for how each subject has affected our lives and for the people God revealed wisdom to as they sought to understand Creation Each content area is thoroughly explored giving students a good foundation in each discipline Semester 1 Math and Physics Numbers surround us Just try to make it through a day without using any It's impossible telephone numbers calendars volume settings shoe sizes speed limits weights street numbers microwave timers TV channels and the list goes on and on The many advancements and branches of mathematics were developed through the centuries as people encountered problems and relied upon math to solve them It's amazing how ten simple digits can be used in an endless number of ways to benefit man The development of these ten digits and their many uses is the fascinating story in Exploring the World of Mathematics Physics is a branch of science that many people consider to be too complicated to understand John Hudson Tiner puts this myth to rest as he explains the fascinating world of physics in a way that students can comprehend Did you know that a feather and a lump of lead will fall at the same rate in a vacuum Learn about the history of physics from Aristotle to Galileo to Isaac Newton to the latest advances Discover how the laws of motion and gravity affect everything from the normal activities of everyday life to launching rockets into space Learn about the effects of inertia first hand during fun and informative experiments Exploring the World of Physics is a great tool for student who want to have a deeper understanding of the important and interesting ways that physics affects our lives Semester 2 Biology and Chemistry The field of biology focuses on living things from the smallest microscopic protozoa to the largest mammal In this book you will read and explore the life of plants insects spiders and other arachnids life in water reptiles birds and mammals highlighting God's amazing creation You will learn about biological classification how seeds spread around the world long term storage of energy how biologists learned how the stomach digested food the plant that gave George de Mestral the idea of Velcro and so much more For most of history biologists used the visible appearance of plants or animals to classify them They grouped plants or animals with similar looking features into families Starting in the 1990 s biologists have extracted DNA and RNA from cells as a guide to how plants or animals should be grouped Like visual structures these reveal the underlying design of creation Exploring the World of Biology is a fascinating look at life from the smallest proteins and spores to the complex life systems of humans and animals Chemistry is an amazing branch of science that affects us every day yet few people realize it or even give it much thought Without chemistry there would be nothing made of plastic there would be no rubber tires no tin cans no televisions no microwave ovens or something as simple as wax paper This book presents an exciting and intriguing tour through the realm of chemistry as each chapter unfolds with facts and stories about the discoveries of discoverers Find out why pure gold is not used for jewelry or coins Join Humphry Davy as he made many chemical discoveries and learn how

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affordances of the room How can teachers address familiar classroom management challenges in these unfamiliar spaces If assessment and rapid feedback are critical in active learning how do they work in a room filled with circular tables and no central focus point How do instructors balance group learning with the needs of the larger class How can students be held accountable when many will necessarily have their backs facing the instructor How can instructors evaluate the effectiveness of their teaching in these spaces This book is intended for faculty preparing to teach in or already working in this new classroom environment for administrators planning to create ALCs or experimenting with provisionally designed rooms and Chemistry Carson-Dellosa for faculty developers helping teachers transition to using these new spaces Publishing, 2015-03-16 Chemistry for grades 9 to 12 is designed to aid in the review and practice of chemistry topics Chemistry covers topics such as metrics and measurements matter atomic structure bonds compounds chemical equations molarity and acids and bases The book includes realistic diagrams and engaging activities to support practice in all areas of chemistry The 100 Series science books span grades 5 to 12 The activities in each book reinforce essential science skill practice in the areas of life science physical science and earth science The books include engaging grade appropriate activities and clear thumbnail answer keys Each book has 128 pages and 100 pages or more of reproducible content to help students review and reinforce essential skills in individual science topics. The series will be aligned to current science standards GATE Question Bank - Biotechnology Mocktime Publication, 2400 MCQs GATE Biotechnology Chapterwise Question Bank Based on New Syllabus Physics Workbook For Dummies Steven Holzner, 2007-10-05 Do you have a handle on basic physics terms and concepts but your problem solving skills could use some static friction Physics Workbook for Dummies helps you build upon what you already know to learn how to solve the most common physics problems with confidence and ease Physics Workbook for Dummies gets the ball rolling with a brief overview of the nuts and bolts i e converting measures counting significant figures applying math skills to physics problems etc before getting into the nitty gritty If you re already a pro on the fundamentals you can skip this section and jump right into the practice problems There you ll get the lowdown on how to take your problem solving skills to a whole new plane without ever feeling like you ve been left spiraling down a black hole With easy to follow instructions and practical tips Physics Workbook for Dummies shows you how to you unleash your inner Einstein to solve hundreds of problems in all facets of physics such as Acceleration distance and time Vectors Force Circular motion Momentum and kinetic energy Rotational kinematics and rotational dynamics Potential and kinetic energy Thermodynamics Electricity and magnetism Complete answer explanations are included for all problems so you can see where you went wrong or right Plus you ll get the inside scoop on the ten most common mistakes people make when solving physics problems and how to avoid them When push comes to shove this friendly guide is just what you need to set your physics problem solving skills in motion Chemistry, 2015-03-16 Chemistry for grades 9 to 12 is designed to aid in the review and practice of chemistry topics Chemistry covers topics such as metrics and measurements

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