

LOTS OF CONVERSIONS!

Name Grove
Period Date / /

8 • Why Do Hot Air Balloons Float?

THE IDEAL GAS LAW

$PV = nRT$ where

P = pressure in atmosphere
V = volume in liters
n = number of moles of gas
R = Universal Gas Constant = 0.0821 L·atm/mol·K
T = Kelvin temperature

See scratch paper

1. How many moles of oxygen will occupy a volume of 2.50 liters at 1.20 atm and ~~25°C~~ ^{298 K}?

$$PV = nRT \quad n = \frac{PV}{RT} = \frac{1.20 \text{ atm} \cdot 2.50 \text{ L}}{0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K} \cdot 298 \text{ K}} = 0.947 \text{ mol}$$
2. What volume will 2.00 moles of nitrogen occupy at ~~720 mmHg~~ ^{0.947 atm} and ~~20°C~~ ^{293 K}?

$$V = \frac{nRT}{P} = \frac{2.00 \text{ mol} \cdot 0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K} \cdot 293 \text{ K}}{0.947 \text{ atm}} = 500 \text{ L}$$
3. What pressure will be exerted by ~~25.0 g~~ ^{0.568 mol} of CO_2 at temperature of ~~25°C~~ ^{298 K} and a volume of ~~500 mL~~ ^{0.500 L}?

$$P = \frac{nRT}{V} = \frac{0.568 \text{ mol} \cdot 0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K} \cdot 298 \text{ K}}{0.500 \text{ L}} = 0.947 \text{ atm}$$
4. At what temperature will ~~5.00 g~~ ^{0.704 mol} of Cl_2 exert a pressure of ~~900 mmHg~~ ^{1.18 atm} at a volume of ~~250 mL~~ ^{0.250 L}?

$$T = \frac{PV}{nR} = \frac{1.18 \text{ atm} \cdot 0.250 \text{ L}}{0.704 \text{ mol} \cdot 0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K}} = 293 \text{ K}$$
5. What is the ~~density~~ ^{9/L} of NH_3 at ~~800 mmHg~~ ^{1.05 atm} and ~~25°C~~ ^{298 K}?

$$\rho = \frac{PM}{RT} = \frac{1.05 \text{ atm} \cdot 17.03 \text{ g/mol}}{0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K} \cdot 298 \text{ K}} = 0.750 \text{ g/L}$$
6. If the density of a gas is 1.2 g/L at ~~745 mmHg~~ ^{0.980 atm} and ~~20°C~~ ^{293 K}, what is its molar mass? $\frac{d \cdot RT}{P} = \frac{1.2 \text{ g/L} \cdot 0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K} \cdot 293 \text{ K}}{0.980 \text{ atm}} = 29.1 \text{ g/mol}$
7. How many moles of nitrogen gas will occupy a volume of ~~347 mL~~ ^{0.347 L} at ~~6680 mmHg~~ ^{8.79 atm} and ~~27°C~~ ^{300 K}?

$$n = \frac{PV}{RT} = \frac{8.79 \text{ atm} \cdot 0.347 \text{ L}}{0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K} \cdot 300 \text{ K}} = 0.120 \text{ mol}$$
8. What volume will ~~454 grams~~ ^{225 mol} (1 lb) of hydrogen occupy at 1.05 atm and ~~25°C~~ ^{298 K}?

$$V = \frac{nRT}{P} = \frac{225 \text{ mol} \cdot 0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K} \cdot 298 \text{ K}}{1.05 \text{ atm}} = 500 \text{ L}$$
9. Find the number of grams of CO_2 that exert a pressure of ~~285 mmHg~~ ^{1.03 atm} at a volume of 32.5 L and a temperature of ~~32°C~~ ^{305 K}?

$$n = \frac{PV}{RT} = \frac{1.03 \text{ atm} \cdot 32.5 \text{ L}}{0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K} \cdot 305 \text{ K}} = 1.33 \text{ mol}$$
10. An elemental gas has a mass of 10.3 g. If the volume is 58.4 L and the pressure is ~~758 mmHg~~ ^{0.997 atm} at a temperature of ~~25°C~~ ^{275.5 K}, what is the gas? $P \cdot V = n \cdot R \cdot T \rightarrow n = \frac{P \cdot V}{R \cdot T} = \frac{0.997 \text{ atm} \cdot 58.4 \text{ L}}{0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K} \cdot 275.5 \text{ K}} = 2.58 \text{ mol}$
 $\text{Molar Mass} = \frac{\text{Mass}}{\text{Moles}} = \frac{10.3 \text{ g}}{2.58 \text{ mol}} = 4.0 \text{ g/mol}$

Ideal Gas Law Answers Pg 24

Christian G. Meyer

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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forensics and more medical related topics Numerous worked examples in the text now include Analysis and Synthesis sections which allow engineers to explore concepts in greater depth and discuss outside relevance *Gaseous Fuels*
American Gas Association, 1948

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