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systems and Wiring diagrams Models covered include Kia Sephia 1994 2001 Kia Spectra 2000 2009 Kia Sportage 2005 2020

Kia Sorento, '03-'13 Haynes Publishing, 2018-05-01 Each Haynes manual provides specific and detailed instructions for performing everything from basic maintenance and troubleshooting to a complete overhaul of the machine in this case the Kia Sorento model years 2003 through 2013 Do it yourselfers will find this service and repair manual more comprehensive than the factory manual making it an indispensable part of their tool box A typical Haynes manual covers general information troubleshooting lubrication and routine maintenance engine top end engine lower end primary drive clutch and external shift mechanism transmission and internal shift mechanism engine management system electrical system wheels tires and drivebelt front suspension and steering rear suspension brakes body and color wiring diagrams An index makes the manual easy to navigate

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for the nitration 4.01 ml of nitric acid 4.00 ml of sulfuric acid and 4.56 ml bromobenzene were used to complete the reaction with these measurements bromobenzene was found to be the limiting reagent since only 4.56 ml of bromobenzene was used the amount of product predicted to be recovered was 8.7533 g

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the aim of the experiment is to determine and study electrophilic aromatic substitution reaction especially nitration of methyl benzoate to give methyl m nitrobenzoate thus determine the weight melting point and percentage yield of the pure methyl m nitrobenzoate abstract

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in this experiment we will carry out the nitration of bromobenzene bromine substituted on an aromatic ring is an ortho para directing group it will direct an incoming electrophile to the ortho and para positions it is also deactivating meaning that its reactions will be slower than the identical reactions of benzene

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section 20711 lab experiment nitration of bromobenzene introduction the chemical reaction of nitration consists of a nitro group being added to or substituted in a molecule nitration can basically be carried out by a mixture of concentrated nitric acid and sulfuric acid this mixture is useful to obtain the active nitronium ion

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in the nitration of bromobenzene a mixture of sulfuric acid and nitric acid react with one another nitric acid the weaker of the two acids will serve as base to form nitronium ion bromobenzene will attack the nitronium ion by nucleophilic attack

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then actual is your actual findings yield actual theoretical x 100 what are the possible side reactions of nitration of bromobenzene dinitration can occur at ortho para or at both ortho positions occurs in overheating meta product happens very rarely because bromine is an ortho para director

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the soap in the lab is not suitable for this purpose 2 return to the lab and apply mineral oil to the affected area theory and explanation of the nitration of bromobenzene if the temperature of the reaction mixture exceeds 60 °C

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the bromine substituent in bromobenzene is an ortho para director the para position is favored due to steric effects in the reaction that takes place in this lab a nitro group is substituted onto bromobenzene at the para position to form 1-bromo-4-nitrobenzene

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over 5 minutes slowly pipet this bromobenzene into the icy cold nitration mixture HNO_3 3 mL H_2SO_4 in the 125 mL flask and stir continue to stir the reaction mixture in the ice bath for 30 minutes do not heat in a new 250 mL erlenmeyer

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on the right the arenium cation has a nitro set with incredibly small 0.6 Å bond distances this will result in an optimization failure as the atoms are even close for the programme to accommodate their electrons demonstration of creating the p arenium cation by nitration of bromobenzene

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the purpose of this lab was to observe where substituents are added on a benzene ring by way of electrophilic aromatic substitution the electrophilic source in the nitration of bromobenzene comes from a nitronium ion NO_2^+ which acts as the electrophilic species

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nitration of bromobenzene is an electrophilic aromatic substitution reaction the electrophile will be formed from 4 0ml of nitric acid and 4 0ml of sulfuric acid and will then react with 4 5ml of bromobenzene this will produce o bromonitrobenzene and p bromonitrobenzene

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