

Chemical Reactions

Section 9.1 Reactions and Equations

pages 282–288

Practice Problems

pages 284–287

Write skeleton equations for the following word equations.

1. Hydrogen and bromine gases react to yield hydrogen bromide.

hydrogen(g) + bromine(g) →
hydrogen bromide(g)



2. When carbon monoxide and oxygen react, carbon dioxide forms.

carbon monoxide(g) + oxygen(g) →
carbon dioxide(g)



3. **Challenge** Write the word equation and the skeleton equation for the following reaction: when heated, solid potassium chlorate yields solid potassium chloride and oxygen gas.

Potassium chlorate(s) →
potassium chloride(s) + oxygen(g)



Write chemical equations for each of the following reactions.

4. In water, iron(III) chloride reacts with sodium hydroxide, producing solid iron(III) hydroxide and sodium chloride.



5. Liquid carbon disulfide reacts with oxygen gas, producing carbon dioxide gas and sulfur dioxide gas.



6. **Challenge** A piece of zinc metal is added to a solution of hydrogen sulfate. This reaction produces a gas and a solution of zinc sulfate.



Section 9.1 Assessment

page 288

7. **Explain** why it is important that a chemical equation be balanced.

Because mass is neither created nor destroyed in chemical reactions, the numbers of atoms of all elements must be equal on both sides of the reaction arrow.

8. **List** three types of physical evidence that indicate a chemical reaction has occurred.

Answers may include release or absorption of energy, change in color, change in odor, formation of a gas, or formation of a solid.

9. **Compare and contrast** a skeleton equation and a chemical equation.

The skeleton equation includes the formulas of reactants and products. The chemical equation gives the relative amounts of reactants and products.

10. **Explain** why it is important to reduce coefficients in a balanced equation to the lowest possible whole-number ratio.

Coefficients in the lowest ratio most clearly indicate the relative amounts of substances in a reaction.

11. **Analyze** When balancing a chemical equation, can you adjust the subscript in a formula? Explain.

No. Doing so changes the identity of the substance.

Mcgraw Hill Chemical Reactions Study Guide

Guillaume Favre



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Table of Contents Mcgraw Hill Chemical Reactions Study Guide

1. Understanding the eBook Mcgraw Hill Chemical Reactions Study Guide
 - The Rise of Digital Reading Mcgraw Hill Chemical Reactions Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Mcgraw Hill Chemical Reactions Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mcgraw Hill Chemical Reactions Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mcgraw Hill Chemical Reactions Study Guide
 - Personalized Recommendations
 - Mcgraw Hill Chemical Reactions Study Guide User Reviews and Ratings
 - Mcgraw Hill Chemical Reactions Study Guide and Bestseller Lists
5. Accessing Mcgraw Hill Chemical Reactions Study Guide Free and Paid eBooks
 - Mcgraw Hill Chemical Reactions Study Guide Public Domain eBooks
 - Mcgraw Hill Chemical Reactions Study Guide eBook Subscription Services
 - Mcgraw Hill Chemical Reactions Study Guide Budget-Friendly Options

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 - Mcgraw Hill Chemical Reactions Study Guide Enhanced eBook Features
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 - Highlighting and Note-Taking Mcgraw Hill Chemical Reactions Study Guide
 - Interactive Elements Mcgraw Hill Chemical Reactions Study Guide
8. Staying Engaged with Mcgraw Hill Chemical Reactions Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mcgraw Hill Chemical Reactions Study Guide
9. Balancing eBooks and Physical Books Mcgraw Hill Chemical Reactions Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mcgraw Hill Chemical Reactions Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Mcgraw Hill Chemical Reactions Study Guide
 - Setting Reading Goals Mcgraw Hill Chemical Reactions Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mcgraw Hill Chemical Reactions Study Guide
 - Fact-Checking eBook Content of Mcgraw Hill Chemical Reactions Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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