



CHEMISTRY

CHEAT SHEET

EQUILIBRIUM

EQUILIBRIUM CHEAT SHEET

Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time having reached Equilibrium. Equilibrium reactions are represented by the symbol \rightleftharpoons .

<p>What is Equilibrium?</p> <p>Equilibrium is a state of balance between two or more opposing processes.</p> <p>In chemistry, it refers to a reaction where the forward and reverse rates are equal, and the concentrations of reactants and products remain constant over time.</p> <p>Equilibrium is reached when the system has no net change in the concentrations of the reactants and products.</p> <p>Equilibrium is a dynamic process, meaning that the reaction is still occurring, but the rates of the forward and reverse reactions are equal.</p>	<p>Equilibrium Constant</p> <p>The equilibrium constant, K, is a measure of the extent to which a reaction proceeds. It is defined as the ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	<p>Graph showing concentration vs. time for a reaction at equilibrium.</p> <p>The concentration of reactants decreases over time, while the concentration of products increases, both reaching a constant value at equilibrium.</p>																								
<p>Equilibrium Constant, K_c</p> <p>The equilibrium constant, K_c, is defined as the ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$ <p>where $[A]$, $[B]$, $[C]$, and $[D]$ are the concentrations of the reactants and products in moles per liter (M).</p>	<table><tr><th>Equilibrium Constant, K_c</th><th>Meaning</th></tr><tr><td>$K_c > 1$</td><td>Products are favored at equilibrium (reaction proceeds to the right)</td></tr><tr><td>$K_c < 1$</td><td>Reactants are favored at equilibrium (reaction proceeds to the left)</td></tr><tr><td>$K_c = 1$</td><td>Reactants and products are at equal concentrations</td></tr></table>	Equilibrium Constant, K_c	Meaning	$K_c > 1$	Products are favored at equilibrium (reaction proceeds to the right)	$K_c < 1$	Reactants are favored at equilibrium (reaction proceeds to the left)	$K_c = 1$	Reactants and products are at equal concentrations																	
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<p>Equilibrium Constant, K_p</p> <p>The equilibrium constant, K_p, is defined as the ratio of the partial pressures of the products to the partial pressures of the reactants, each raised to the power of its stoichiometric coefficient.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K_p = \frac{(P_C)^c (P_D)^d}{(P_A)^a (P_B)^b}$ <p>where P_A, P_B, P_C, and P_D are the partial pressures of the reactants and products in atmospheres (atm).</p>	<table><tr><th>Equilibrium Constant, K_p</th><th>Meaning</th></tr><tr><td>$K_p > 1$</td><td>Products are favored at equilibrium (reaction proceeds to the right)</td></tr><tr><td>$K_p < 1$</td><td>Reactants are favored at equilibrium (reaction proceeds to the left)</td></tr><tr><td>$K_p = 1$</td><td>Reactants and products are at equal partial pressures</td></tr></table>	Equilibrium Constant, K_p	Meaning	$K_p > 1$	Products are favored at equilibrium (reaction proceeds to the right)	$K_p < 1$	Reactants are favored at equilibrium (reaction proceeds to the left)	$K_p = 1$	Reactants and products are at equal partial pressures																	
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<p>Le Chatelier's Principle</p> <p>Le Chatelier's Principle states that if a system at equilibrium is subjected to a change in concentration, pressure, or temperature, the system will shift its equilibrium position to counteract the change.</p> <p>For example, if the concentration of a reactant is increased, the system will shift to the right, favoring the products, to reduce the concentration of the reactant.</p>	<table><tr><th>Le Chatelier's Principle</th><th>Meaning</th></tr><tr><td>Concentration</td><td>Increasing the concentration of a reactant will shift the equilibrium to the right, favoring the products.</td></tr><tr><td>Pressure</td><td>Increasing the pressure will shift the equilibrium to the side with the fewest moles of gas.</td></tr><tr><td>Temperature</td><td>Increasing the temperature will shift the equilibrium to the side that is endothermic (absorbs heat).</td></tr></table>	Le Chatelier's Principle	Meaning	Concentration	Increasing the concentration of a reactant will shift the equilibrium to the right, favoring the products.	Pressure	Increasing the pressure will shift the equilibrium to the side with the fewest moles of gas.	Temperature	Increasing the temperature will shift the equilibrium to the side that is endothermic (absorbs heat).																	
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Mike Jess

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