



A thermometer is secured with a thermometer clamp to guard against breakage.

Experiment 14

Molar Mass of a Solid

- To observe and measure the effect of a solute on the freezing point of a solvent
- To determine the molar mass (molecular weight) of a nonvolatile, nonelectrolyte solute

OBJECTIVES

The following techniques are used in the Experimental Procedure:



TECHNIQUES

A pure liquid, such as water or ethanol, has characteristic physical properties: the melting point, boiling point, density, vapor pressure, viscosity, surface tension, and additional data listed in handbooks of chemistry. The addition of a **soluble** solute to the liquid forms a homogeneous mixture called a **solution**. The solvent of the solution assumes physical properties that are no longer definite but dependent on the amount of solute added. The vapor pressure of the solvent decreases, the freezing point of the solvent decreases, the boiling point of the solvent increases, and the osmotic pressure of the solvent increases. The degree of the change depends on the *number* of solute particles that have dissolved, *not* on the chemical identity of the solute. These four physical properties that depend on the number of solute particles dissolved in a solvent are called **colligative properties**.

For example, one mole of glucose or urea (neither of which dissociates in water) lowers the freezing point of one kilogram of water by 1.86°C; whereas one mole of sodium chloride lowers the freezing point of one kilogram of water by nearly twice that amount (~3.72°C) because, when dissolved in water, it dissociates into Na^+ and Cl^- providing twice as many moles of solute particles per mole of solute as do glucose or urea.

When freezing ice cream at home, a salt-ice water mixture provides a lower temperature bath than an ice/water mixture alone. Antifreeze (ethylene glycol, Figure 14.1) added to the cooling system of an automobile reduces the probability of freeze-up in the winter and boiling over in the summer because the antifreeze/water solution has a lower freezing point and a higher boiling point than pure water.

These changes in the properties of pure water that result from the presence of a **nonvolatile solute** are portrayed by the phase diagram in Figure 14.2, page 190, a plot of vapor pressure versus temperature. The solid lines refer to the equilibrium conditions between the respective phases for pure water; the dashed lines represent the same conditions for an aqueous solution.

INTRODUCTION



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Figure 14.1 Ethylene glycol is a major component of most antifreeze solutions.

Colligative properties: properties of a solvent that result from the presence of the number of solute particles in the solution and not their chemical composition

Nonvolatile solute: a solute that does not have a measurable vapor pressure

J A Beran Lab Manual Experiment 14

Michael Hoelscher



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Laboratory Manual for Principles of General Chemistry J. A. Beran, Mark Lassiter, 2022-08-16 Laboratory Manual for Principles of General Chemistry 11th Edition covers two semesters of a general chemistry laboratory program The material focuses on the lab experiences that reinforce the concepts that not all experimental conclusions are the same and depend on identifying an appropriate experimental procedure selecting the proper apparatus employing the proper techniques systematically analyzing and interpreting the data and minimizing inherent variables As a result of good data a scientific and analytical conclusion is made which may or may not be right but is certainly consistent with the data Experiments write textbooks textbooks don't write experiments A student's scientific literacy grows when experiences and observations associated with the scientific method are encountered Further experimentation provides additional cause effect observations leading to an even better understanding of the experiment The 11th edition's experiments are informative and challenging while offering a solid foundation for technique safety and experimental procedure The reporting and analysis of the data and the pre and post lab questions focus on the intuitiveness of the experiment The experiments may accompany any general chemistry textbook and are compiled at the beginning of each curricular unit An Additional Notes column is included in each experiment's Report Sheet to provide a space for recording observations and data during the experiment Continued emphasis on handling data is supported by the Data Analysis section Laboratory Manual for Principles of General Chemistry Jo Allan Beran, 2010-11-01 This new edition of the Beran lab manual emphasizes chemical principles as well as techniques The manual helps students understand the timing and situations for the various techniques The Beran lab manual has long been a market leading lab manual for general chemistry Each experiment is presented with concise objectives a comprehensive list of techniques and detailed lab intros and step by step procedures Favorite Demonstrations for College Science Brian Shmaefsky, 2004 The book is an all in one compilation of 36 popular classroom demonstrations published since 1993 in the Favorite Demonstration column of NSTA's Journal of College Science Teaching The collection begins with a chapter on safety The Rules of Research From there chapters emphasize conveying scientific principles while making them memorable Английский язык для изучающих химию (B1 - B2). Учебник для вузов Татьяна Барановская, Анна Захарова, Татьяна Поспелова, Юлия Суворова, 2023-11-07 1 2 04 03 01 1 2 CEFR **Английский язык для изучающих химию (B1 - B2). Учебник для СПО** Татьяна Барановская, Анна Захарова, Татьяна Поспелова, Юлия Суворова, 2023-12-01 Wordstool Scientific and Technical Books and Serials in Print, 1984 *Paperbound Books in Print* 1995 Reed Reference Publishing, R5ference Reed, 1995-12 *Proceedings of the National Academy of Sciences of the United States of America* National Academy of Sciences (U.S.), 1989 **Forthcoming Books** Rose Arny, 1996-06 Books in Print Supplement, 1984 **Cumulated Index Medicus**, 1995 **Monthly Index of Russian Accessions** Library of Congress. Processing Department, 1966 **Monthly List of Russian Accessions** Library of Congress. Processing

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