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Nozzle Heat Seal Service Removal Procedure

■ IMPORTANT

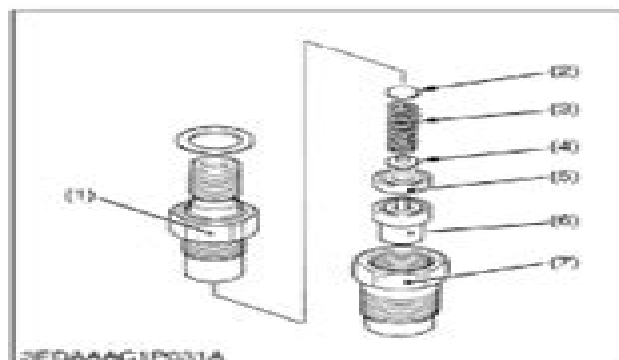
- Use a plus (Phillips head) screw driver that has a Dia. which is bigger than the heat seal hole (Approx. 6 mm (1/4 in.)).
- Drive screw driver lightly into the heat seal hole.
 - Turn screw driver three or four times each way.
 - While turning the screw driver, slowly pull the heat seal out together with the injection nozzle gasket.

If the heat seal drops, repeat the above procedure. Heat seal and injection nozzle gasket must be changed when the injection nozzle is removed for cleaning or for service.

- (1) Plus Screw Driver
(2) Injection Nozzle

- (3) Injection Nozzle Gasket
(4) Heat Seal

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Nozzle Holder

- Secure the nozzle retaining nut (7) with a vise.
- Remove the nozzle holder (1), and take out parts inside.

(When reassembling)

- Assemble the nozzle in clean fuel oil.
- Install the push rod (4), noting its direction.
- After assembling the nozzle, be sure to adjust the fuel injection pressure.

Tightening torque	Nozzle holder	35 to 39 N·m 3.5 to 4.0 kgf·m 25 to 28 lbf·ft
	Nozzle holder assembly	49 to 65 N·m 5.0 to 7.0 kgf·m 37 to 50 lbf·ft

- (1) Nozzle Holder
(2) Adjusting Washer
(3) Nozzle Spring
(4) Push Rod

- (5) Distance Piece
(6) Nozzle Piece
(7) Nozzle Retaining Nut

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Cylinder Head Cover

- Remove the cylinder head cover (1).

(When reassembling)

- After reassembling the head cover, adjust decompression device.

Tightening torque	Head cover mounting nut	6.8 to 10 N·m 1.0 to 1.1 kgf·m 7.3 to 7.9 lbf·ft
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- (1) Cylinder Head Cover

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Kubota Dynamo Manual

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Table of Contents Kubota Dynamo Manual

1. Understanding the eBook Kubota Dynamo Manual
 - The Rise of Digital Reading Kubota Dynamo Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Kubota Dynamo Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Kubota Dynamo Manual
 - User-Friendly Interface
4. Exploring eBook Recommendations from Kubota Dynamo Manual
 - Personalized Recommendations
 - Kubota Dynamo Manual User Reviews and Ratings
 - Kubota Dynamo Manual and Bestseller Lists

5. Accessing Kubota Dynamo Manual Free and Paid eBooks
 - Kubota Dynamo Manual Public Domain eBooks
 - Kubota Dynamo Manual eBook Subscription Services
 - Kubota Dynamo Manual Budget-Friendly Options
6. Navigating Kubota Dynamo Manual eBook Formats
 - ePub, PDF, MOBI, and More
 - Kubota Dynamo Manual Compatibility with Devices
 - Kubota Dynamo Manual Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Kubota Dynamo Manual
 - Highlighting and Note-Taking Kubota Dynamo Manual
 - Interactive Elements Kubota Dynamo Manual
8. Staying Engaged with Kubota Dynamo Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Kubota Dynamo Manual
9. Balancing eBooks and Physical Books Kubota Dynamo Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Kubota Dynamo Manual
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Kubota Dynamo Manual
 - Setting Reading Goals Kubota Dynamo Manual
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Kubota Dynamo Manual
 - Fact-Checking eBook Content of Kubota Dynamo Manual
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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web $m \times c \times \Delta t$ equation 1 where m is the total mass of the solution solute plus solvent c is the specific heat of the solution and Δt is the observed temperature change the specific heat of the solution is generally assumed to be the same as that of
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web the heat of reaction ΔH_{soln} is written after the products in units of kJ mol rxn for example sodium hydroxide dissolves exothermically releasing 44.2 kilojoules per mole dissolved this process is represented as $\text{NaOH (s)} \rightarrow \text{NaOH (aq)}$ $\Delta H_{\text{soln}} = -44.2 \text{ kJ}$

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web in this lab you are challenged to use chemistry to determine which ionic compound makes the best hand warmer in this experiment you will collect data that will allow you to calculate the heat of solution ΔH_{soln} or the amount of energy released or absorbed in kJ/mol of the substance being dissolved

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