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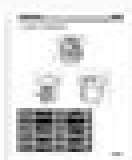
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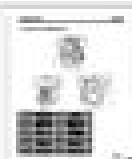
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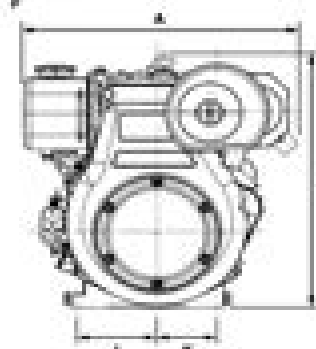
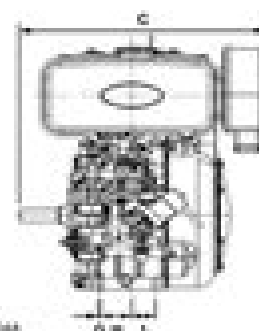
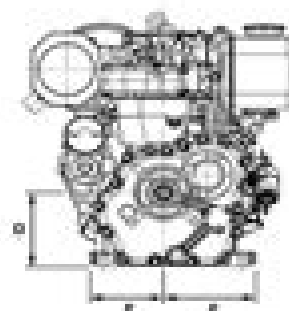


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OC60-E2, OC60-E3, WSM

DIMENSIONS

## ■ K TYPE (WITH STOP SOLENOID) (OC60-E2)



W6CAABP0044

	STANDARD TYPE	K TYPE
A	500 mm (19.69 in.)	500 mm (19.69 in.)
B	501 mm (19.73 in.)	501 mm (19.73 in.)
C	452 mm (17.80 in.)	452 mm (17.80 in.)
D	145 mm (5.71 in.)	145 mm (5.71 in.)
E	130 mm (5.12 in.)	130 mm (5.12 in.)
F	170 mm (6.69 in.)	170 mm (6.69 in.)
G	3 mm (0.12 in.)	3 mm (0.12 in.)
H	57 mm (2.24 in.)	57 mm (2.24 in.)
I	43 mm (1.69 in.)	43 mm (1.69 in.)
J	145 mm (5.71 in.)	145 mm (5.71 in.)
K	110 mm (4.33 in.)	110 mm (4.33 in.)

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# Kubota Oc60 E2 Oc95 E2 Diesel Engine Digital Workshop Repair Manual

**LL Leslie**



## **Kubota Oc60 E2 Oc95 E2 Diesel Engine Digital Workshop Repair Manual:**

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0 2885 its mole percent would be 28 85

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web since the mole fraction of the  $O_2$  is 0.050 it contributes 0.050 or if you will 5.0 of the total pressure thus total pressure is 0.210.050 = 4.2 atm problem 9 a sample of 1.43 g of helium and an unweighed quantity of  $O_2$  are mixed in a flask at room temperature

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