

## Student Notebook Navigator

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## Experiment 1: Mass Volume and Density ✓

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- 1
6. (T) In this experiment, why is the mass of the burette segment with water measured twice for each sample? (hint: What potential error or uncertainty is being accounted for?) (1-2 sentences)
- +
7. (T) Why should the level of water in the buret segment be positioned below the level of water in the 100-mL beaker? (1 phrase/sentence)
- +
8. (T) Why do we move the hot plate and beaker and not the buret segment once the temperature reaches 60 degrees? (1 phrase/sentence)
- +
9. (T) What are the SI units for density? (units only)
- +
10. (B) The following data represents 3 series of measurements of the density of water. The literature value for the density of water is 1.00 g/mL. Use Excel to calculate the mean, standard deviation and CV of each data set and copy the values into the table provided, then answer the questions below. (find the formula for CV in the Lab Manual)

Data A (g/mL)	Data B (g/mL)	Data C (g/mL)
1.01	1.00	1.00
1.05	1.00	1.10
1.00	1.01	0.99
1.05	1.00	1.10
1.00	1.01	1.00

	Data A (g/mL)	Data B (g/mL)	Data C (g/mL)
Mean			
Standard Deviation			
CV			

- a. (T) Which data set is the most precise? (A, B, or C)

+

- a. (T) Which data set is the most accurate? (A, B, or C)

# Hayden Mcneil Chem 115 Lab Manual Answers

**UNC Greensboro Staff, Unc Greensboro**

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the text in the light blue bar.

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