

HISTORY OF THE ATOM

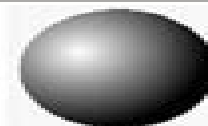
GENERAL CHEMISTRY 1

DATE: SEPT 15, 2020

MADE BY: ELLA SABIDO 11SP

DEMOCRITUS (400 B.C.)

- formulated the term 'atom' and believed that matter can be broken down to its simplest and smallest substance called atom.



ARISTOTLE (300 B.C. - 1800's)

- dismissed the idea of atoms and declared that all matter was made of only four elements: fire, air, water and earth.
- believed that there are only five elements that exist: fire, water, earth and air.



DALTON (1803)

- he theorized that all matter is made of atoms. Atoms are too small to see, "uncuttable," and indestructible.
- developed the atomic theory: All atoms of a given element are exactly alike and atoms of different elements are different.



BILLIARD BALL ATOM

THOMSON (1897)

- he discovered the presence of the electron.
- he proposed a model of the atom that is sometimes called the "Plum Pudding" model. His theory was that atoms are made from a positively charged substance with negatively charged electrons scattered about, like raisins in a pudding or chocolate chips in a cookie.



PLUM PUDDING

RUTHERFORD (1908)

- he performed an experiment using positively charged particles fired at gold foil.
- he theorized that atoms have a small, dense, positively charged center, which he called the "nucleus".



NUCLEUS MODEL

BOHR (1913)

- he suggested that electrons orbit in different levels.
- he studied the structure of atoms based on Rutherford's discoveries.
- the more energy an atom gives off, the closer it is to the nucleus, the more it absorbs, the farther away.



BOHR MODEL OR THE QUANTUM MODEL

MODERN CLOUD THEORY (20th Century)

- according to today's atomic theory, in the electron cloud, electrons whirl around the nucleus billions of times in one second. They are not moving around in random patterns; an electron's location depends upon how much energy the electron has.



COTTON BALL OR MODERN ATOM MODEL



History Of The Atom Notes Key

Vincent Kiernan



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