



At the end of this chapter, you will be able to describe the properties and uses of metals. You will see that the properties of metals are mainly determined by their structure. This will also understand how these properties are used to make products for many different purposes.

You will learn how metals are used to make many different products in the industry of metals and how they are used to make many different products in the industry of metals.

The properties of metals can be determined by the type of their structure. Metals with different structures have different properties. These properties can be used to make many different products.

Finally, you will learn how metals are used to make many different products in the industry of metals and how they are used to make many different products in the industry of metals.

Key knowledge

- The common properties of metals include: malleability, ductility, melting point, heat conductivity, and electrical conductivity. These properties are determined by the nature of metallic bonding and the structure of metals. (4.1)
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4.1 Metallic properties and bonding

More than 90% of the elements in the periodic table are metals. (Figure 4.1.1)

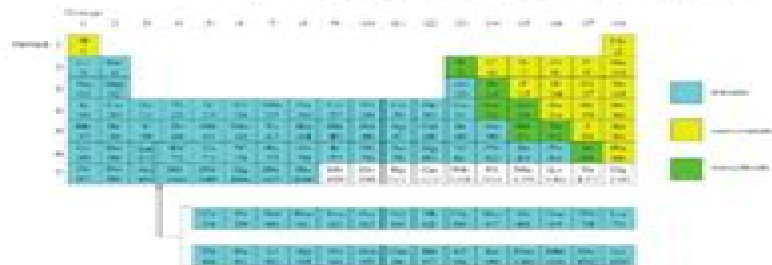


Figure 4.1.1 The distribution of metals, non-metals, and metalloids in the periodic table.

The development of metals in the periodic table is determined by the way the atoms are arranged. The elements in the periodic table are arranged in groups and periods. The elements in the same group have similar properties, and the elements in the same period have similar properties.

Metals are characterized by their properties. They are malleable, ductile, and have a high melting point. They are also good conductors of heat and electricity. These properties are determined by the way the atoms are arranged in the periodic table.

The properties of metals are determined by the way the atoms are arranged in the periodic table. The elements in the same group have similar properties, and the elements in the same period have similar properties. These properties are determined by the way the atoms are arranged in the periodic table.

Table 4.1.1 Properties and uses of some metals

Metals	Properties	Uses
Aluminum	Light, strong, and corrosion-resistant. It is a good conductor of heat and electricity.	Used in aircraft, ships, and buildings. It is also used in the production of aluminum foil and cans.
Iron	Strong and hard. It is a good conductor of heat and electricity.	Used in construction, machinery, and tools. It is also used in the production of steel and cast iron.
Copper	Good conductor of heat and electricity. It is resistant to corrosion.	Used in electrical wiring, plumbing, and coinage. It is also used in the production of brass and bronze.
Gold	Soft and malleable. It is a good conductor of heat and electricity.	Used in jewelry, electronics, and dentistry. It is also used in the production of gold leaf and gold plating.

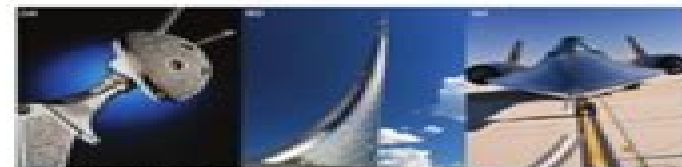


Figure 4.1.2 The use of metals in different applications. (a) A metal wire. (b) A metal pipe. (c) A metal structure.

The properties of metals are determined by the way the atoms are arranged in the periodic table. The elements in the same group have similar properties, and the elements in the same period have similar properties. These properties are determined by the way the atoms are arranged in the periodic table.

GENERAL PROPERTIES OF METALS

Table 4.1.1 shows the properties of some metals and their uses. The properties of metals are determined by the way the atoms are arranged in the periodic table.

- Metals are good conductors of heat and electricity.
- Metals are malleable and ductile.
- Metals have high melting points.
- Metals are resistant to corrosion.

Metals are characterized by their properties. They are malleable, ductile, and have a high melting point. They are also good conductors of heat and electricity. These properties are determined by the way the atoms are arranged in the periodic table.

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