

Chapter 8 Chemical Reactions Chapter Review Answers

1. The evolution of energy, the production of a gas, the formation of a precipitate, and a color change.
2. A correctly written chemical equation must represent the known facts, contain the correct formulas for the reactants and products, and satisfy the law of conservation of mass.
3.
 - a. It is a number that appears in front of a formula in a chemical equation.
 - b. The number of atoms of each type in the formula is multiplied by the coefficient.
4. methane + oxygen \rightarrow carbon dioxide + water;
 $\text{CH}_4(\text{g}) + \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g})$;
 $\text{CH}_4(\text{g}) + 2\text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{g})$
5. A chemical equation provides the number of molecules, moles, and atoms of each element or compound composing the reactants and products as well as the mole ratios of the reactants to each other and to the products.
6. Word and formula equations have only qualitative meaning; they do not express the relative quantities of a reaction's reactants or products.
7.
 - a. a solution in water
 - b. a substance that accelerates a chemical reaction but can be recovered unchanged
 - c. a chemical reaction in which the products re-form the original reactants
8.
 - a. KOH
 - b. $\text{Ca}(\text{NO}_3)_2$
 - c. Na_2CO_3
 - d. CCl_4
 - e. MgBr_2
9. (1) Balance different types of atoms one at a time; (2) balance types of atoms that appear only once on each side of the equation first; (3) balance as single units any polyatomic ions that appear on both sides of the equation; and (4) balance H atoms and O atoms last.
10.
 - a. 6N
 - b. 4H, 2O
 - c. 4H, 4N, 12O
 - d. 2Ca, 4O, 4H
 - e. 3Ba, 6Cl, 18O
 - f. 5Fe, 10N, 30O
 - g. 12Mg, 8P, 32O
 - h. 4N, 16H, 2S, 8O
 - i. 12Al, 188Se, 72O
 - j. 12C, 32H
11.
 - a. $2\text{ZnS}(\text{s}) + 3\text{O}_2(\text{g}) \rightarrow 2\text{ZnO}(\text{s}) + 2\text{SO}_2(\text{g})$
 - b. $2\text{HCl}(\text{aq}) + \text{Ba}(\text{OH})_2(\text{aq}) \rightarrow \text{BaCl}_2(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$
 - c. $2\text{HNO}_3(\text{aq}) + \text{Ca}(\text{OH})_2(\text{aq}) \rightarrow \text{Ca}(\text{NO}_3)_2(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$
12.
 - a. Solid zinc sulfide reacts with oxygen gas to form solid zinc oxide and sulfur dioxide gas.
 - b. When solid calcium hydride is added to water, aqueous calcium hydroxide and hydrogen gas are formed.
 - c. Aqueous silver nitrate mixed with aqueous potassium iodide produces a precipitate of silver iodide and aqueous potassium nitrate.
13.
 - a. $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$
 - b. $2\text{Al} + \text{Fe}_2\text{O}_3 \rightarrow \text{Al}_2\text{O}_3 + 2\text{Fe}$
 - c. $\text{Pb}(\text{CH}_3\text{COO})_2 + \text{H}_2\text{S} \rightarrow \text{PbS} + 2\text{CH}_3\text{COOH}$
14.
 - a. LiO_2 is an incorrect formula; $4\text{Li} + \text{O}_2 \rightarrow 2\text{Li}_2\text{O}$
 - b. H_2Cl_2 is an incorrect formula; $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$
 - c. MgO_2 is an incorrect formula, and the equation as written is not balanced; $\text{MgCO}_3 \rightarrow \text{MgO} + \text{CO}_2$
 - d. I is an incorrect formula for iodine, and the equation is not balanced; $2\text{NaI} + \text{Cl}_2 \rightarrow 2\text{NaCl} + \text{I}_2$
15.
 - a. $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$
 - b. $\text{P}_2\text{O}_{10} + 6\text{H}_2\text{O} \rightarrow 4\text{H}_3\text{PO}_4$
 - c. $\text{Fe}_2\text{O}_3 + 3\text{CO} \rightarrow 2\text{Fe} + 3\text{CO}_2$
16. $\text{CH}_4(\text{g}) + 4\text{Cl}_2(\text{g}) \rightarrow \text{CCl}_4(\text{l}) + 4\text{HCl}(\text{g})$
17.
 - a. $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$
 - b. $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$
 - c. $2\text{Li} + \text{Cl}_2 \rightarrow 2\text{LiCl}$
 - d. $\text{Ca} + \text{I}_2 \rightarrow \text{CaI}_2$
18.
 - (1) In a synthesis reaction, two or more substances combine to form a new substance: $\text{A} + \text{X} \rightarrow \text{AX}$
 - (2) In a decomposition reaction, a single compound undergoes a reaction that produces two or more simpler substances: $\text{AX} \rightarrow \text{A} + \text{X}$
 - (3) In a single-displacement reaction, one element replaces another element in a compound:
 $\text{A} + \text{BX} \rightarrow \text{AX} + \text{B}$ or $\text{Y} + \text{RX} \rightarrow \text{RY} + \text{X}$

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Williams & Wilkins Review



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Solutions Guide for Introductory Chemistry James F. Hall, 1996 **Complete Preparation for the MCAT** Williams & Wilkins Review, 1998-04 Here is the most respected test prep book for the Medical College Admission Test you can buy featuring an active learning approach for a better understanding of the exam's content and a better chance for success. Unique to this guide are coverage of all recent changes in the MCAT plus a step by step plan for sharpening cognitive skills developing problem solving skills and critical thinking. This thorough guide replaces expensive test preparation courses while giving students exactly what they need to get ready for the MCAT.

Physical Chemistry of Metallurgical Processes, Second Edition Mohammad Shamsuddin, 2021-06-22 This updated second edition retains its classroom tested treatment of physical chemistry of metallurgical topics such as roasting of sulfide minerals matte smelting converting structure properties and theories of slag reduction of oxides and reduction smelting interfacial phenomena steelmaking secondary steelmaking role of halides in extraction of metals refining hydrometallurgy and electrometallurgy and adds new data in worked out examples as well as up to date references to the literature. The book further explains the physical chemistry of various metallurgical topics steps involved in extraction of metals such as roasting matte smelting converting reduction smelting steelmaking reactions deoxidation stainless steelmaking vacuum degassing refining leaching chemical precipitation ion exchange solvent extraction cementation gaseous reduction and electrowinning. Each topic is illustrated with appropriate examples of applications of the technique in extraction of some common reactive rare or refractory metal together with worked out problems explaining the principle of the operation. The problems require imagination and critical analyses and also encourage readers for creative application of thermodynamic data in metal extraction. Updates and condenses text throughout the book by sequential arrangement of paragraphs in different chapters. Maximizes readers understanding of the physicochemical principles involved in extraction production of common and rare reactive metals by pyro as well as hydrometallurgical routes. Reinforces concepts presented with worked examples in each chapter explaining the process steps. Explains the physical chemistry of various metallurgical steps such as roasting matte smelting converting and reduction smelting steelmaking aqueous processing etc in extraction of metals. Collects and uniformly presents scattered information on physicochemical principles of metal production from various books and journals.

Bulletin of the Atomic Scientists, 1958-01 The Bulletin of the Atomic Scientists is the premier public resource on scientific and technological developments that impact global security. Founded by Manhattan Project Scientists the Bulletin's iconic Doomsday Clock stimulates solutions for a safer world.

Handbook of Systems Biology Marian Walhout, Marc Vidal, Job Dekker, 2012-12-31 This book provides an entry point into Systems Biology for researchers in genetics molecular biology cell biology microbiology and biomedical science to understand the key concepts to expanding their work. Chapters organized around broader themes of Organelles and Organisms Systems Properties of Biological Processes Cellular Networks and Systems Biology and Disease.

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Flood Handbook Saeid Eslamian,Faezeh A. Eslamian,2022-04-03 Floods are difficult to prevent but can be managed in order to reduce their environmental social cultural and economic impacts Flooding poses a serious threat to life and property and therefore it s very important that flood risks be taken into account during any planning process This handbook presents different aspects of flooding in the context of a changing climate and across various geographical locations Written by experts from around the world it examines flooding in various climates and landscapes taking into account environmental ecological hydrological and geomorphic factors and considers urban agriculture rangeland forest coastal and desert areas Features Presents the main principles and applications of the science of floods including engineering and technology natural science as well as sociological implications Examines flooding in various climates and diverse landscapes taking into account environmental ecological hydrological and geomorphic factors Considers floods in urban agriculture rangeland forest coastal and desert areas Covers flood control structures as well as preparedness and response methods Written in a global context by contributors from around the world *Practical Chemistry* Lyman Churchill Newell,1922 Technical Book Review Index ,1946 **Descriptive Inorganic Chemistry** James E. House,Kathleen A. House,2015-09-10 House s Descriptive Inorganic Chemistry Third Edition provides thoroughly updated coverage of the synthesis reactions and properties of elements and inorganic compounds Ideal for the one semester ACS recommended sophomore or junior level course in descriptive inorganic chemistry this resource offers a readable and engaging survey of the broad spectrum of topics that deal with the preparation properties and use of inorganic materials Using rich graphics to enhance content and maximize learning the book covers the chemical behavior of the elements acid base chemistry coordination chemistry organometallic compounds and numerous other topics to provide a coherent treatment of the field The book pays special attention to key

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