

Limiting Reagents and Percentage Yield Worksheet

1. Consider the reaction



- a) 80.0 grams of iodine(V) oxide, I_2O_5 , reacts with 28.0 grams of carbon monoxide, CO . Determine the mass of iodine I_2 , which could be produced?

80 g I_2O_5	1 mol I_2O_5	1 mol I_2		XS
1	333.8 g I_2O_5	1 mol I_2O_5		
28 g CO	1 mol CO	1 mol I_2	253.8 g I_2	
1	28 g CO	5 mol CO	1 mol I_2	= 50.8 g I_2

- b) If, in the above situation, only 0.160 moles, of iodine, I_2 was produced.
i) what mass of iodine was produced?

0.160 mol I_2	253.8 g I_2	= 40.6 g I_2
	1 mol I_2	

- ii) what percentage yield of iodine was produced.

40.6 g I_2	$\times 100 =$	79.9 %
50.8 g I_2		

2. Zinc and sulphur react to form zinc sulphide according to the equation.



If 25.0 g of zinc and 30.0 g of sulphur are mixed,

- a) Which chemical is the limiting reactant?

25.0 g Zn	1 mol Zn	1 mol ZnS	= 382 mol ZnS	LR
1	65.41 g Zn	1 mol Zn		
30.0 g S	1 mol S	1 mol ZnS	= 934 mol ZnS	XS
1	32.1 g S	1 mol S		

- b) How many grams of ZnS will be formed?

382 mol ZnS	97.51 g ZnS	= 37.2 g ZnS
	1 mol ZnS	

- c) How many grams of the excess reactant will remain after the reaction is over?

$$.934 \text{ mol ZnS} - .382 \text{ mol ZnS} = .552 \text{ mol ZnS}$$

.552 mol ZnS	1 mol S	32.1 g S	= 17.7 g S
	1 mol ZnS	1 mol S	

3.

Limiting Reactant And Percent Yield Answers

Kenneth W. Raymond



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Table of Contents Limiting Reactant And Percent Yield Answers

1. Understanding the eBook Limiting Reactant And Percent Yield Answers
 - The Rise of Digital Reading Limiting Reactant And Percent Yield Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Limiting Reactant And Percent Yield Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Limiting Reactant And Percent Yield Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Limiting Reactant And Percent Yield Answers
 - Personalized Recommendations
 - Limiting Reactant And Percent Yield Answers User Reviews and Ratings
 - Limiting Reactant And Percent Yield Answers and Bestseller Lists
5. Accessing Limiting Reactant And Percent Yield Answers Free and Paid eBooks

- Limiting Reactant And Percent Yield Answers Public Domain eBooks
 - Limiting Reactant And Percent Yield Answers eBook Subscription Services
 - Limiting Reactant And Percent Yield Answers Budget-Friendly Options
6. Navigating Limiting Reactant And Percent Yield Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Limiting Reactant And Percent Yield Answers Compatibility with Devices
 - Limiting Reactant And Percent Yield Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Limiting Reactant And Percent Yield Answers
 - Highlighting and Note-Taking Limiting Reactant And Percent Yield Answers
 - Interactive Elements Limiting Reactant And Percent Yield Answers
 8. Staying Engaged with Limiting Reactant And Percent Yield Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Limiting Reactant And Percent Yield Answers
 9. Balancing eBooks and Physical Books Limiting Reactant And Percent Yield Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Limiting Reactant And Percent Yield Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Limiting Reactant And Percent Yield Answers
 - Setting Reading Goals Limiting Reactant And Percent Yield Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Limiting Reactant And Percent Yield Answers
 - Fact-Checking eBook Content of Limiting Reactant And Percent Yield Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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