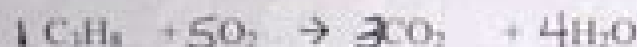


## Limiting Reagent Worksheet #1

1. Given the following reaction: (Balance the equation first!)



a) If you start with 14.8 g of  $\text{C}_3\text{H}_8$  and 3.44 g of  $\text{O}_2$ , how many grams of  $\text{H}_2\text{O}$  are produced?

$$\frac{14.8 \text{ g C}_3\text{H}_8}{44.097 \text{ g C}_3\text{H}_8} \times \frac{1 \text{ mol C}_3\text{H}_8}{1 \text{ mol C}_3\text{H}_8} \times \frac{4 \text{ mol H}_2\text{O}}{1 \text{ mol C}_3\text{H}_8} \times \frac{18.015 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = \frac{1066.488}{44.097} = 24.19 \text{ g}$$

$$\frac{3.44 \text{ g O}_2}{31.998 \text{ g O}_2} \times \frac{1 \text{ mol O}_2}{1 \text{ mol O}_2} \times \frac{4 \text{ mol H}_2\text{O}}{5 \text{ mol O}_2} \times \frac{18.015 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = \frac{247.8564}{159.99} = 1.549 \text{ g}$$

b) Which is the limiting reagent (which reactant created less  $\text{H}_2\text{O}$ )? \_\_\_\_\_

How many grams of the non-limiting reagent are actually used? \_\_\_\_\_

How many grams of the non-limiting reagent are in excess (left over)? \_\_\_\_\_

Following equation:



# Limiting Reactants Study Guide Answer Key

**Tobias Bleicker**



## Limiting Reactants Study Guide Answer Key:

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