



ANSWER GUIDE
PHYSICAL SCIENCES: GRADE 12
CLASS TEST: ACIDS & BASES

Written on: 25.06.2020

Total Marks: 65

1	1.1	B ✓✓	2
	1.2	C ✓✓	2
2	Hydrochloric acid is a strong acid and sodium hydroxide is a strong base, ✓ therefore the equivalence point will be reached at a pH close to 7, ✓ which is in the colour change range of bromothymol blue.		2
3	3.1	An acid is a proton donor. ✓	1
	3.2	A concentrated acid is an acid solution with a high proportion of acid (solute) to water (solvent), ✓✓ regardless of the degree of ionisation or dissociation of the acid in the solution. A strong acid is an acid that dissociates or ionises almost completely when in solution, regardless of its concentration. ✓✓	4
4	4.1	LiOH ✓	5
	4.2	H ₃ PO ₄ ✓	
	4.3	HNO ₃ ✓	
	4.4	HCOOH ✓	
	4.5	NaHCO ₃ ✓	
5	$\text{H}_2\text{SO}_4 + 2\text{H}_2\text{O} \checkmark \rightarrow 2\text{H}_3\text{O}^+ \checkmark + \text{SO}_4^{2-} \checkmark$ OR: $\text{H}_2\text{SO}_4 + \text{H}_2\text{O} \rightarrow \text{H}_3\text{O}^+ + \text{HSO}_4^-$		3
6	6.1	A standard solution is a solution of known concentration. ✓✓	2
	6.2	$c = \frac{m}{MV} \quad \therefore m = cMV \checkmark$ $= (0,15)(106)(0,25) \checkmark$ $= \underline{3,975 \text{ g Na}_2\text{CO}_3} \checkmark$	3
	6.3	$\frac{n_a}{n_b} = \frac{c_a V_a}{c_b V_b} \quad \therefore c_a = \frac{n_a c_b V_b}{n_b V_a} \checkmark$ $= \frac{(1)(0,15)(25)}{(1)(27,5)} \checkmark$ $= \underline{1,38 \times 10^{-1} \text{ mol} \cdot \text{dm}^{-3}} \checkmark$	4
7	pH = -log[H ⁺] ✓ and: HCl dissociates completely $\therefore 0,2 \text{ mol} \cdot \text{dm}^{-3} \text{ H}^+$ ions = -log[0,2] ✓ = 0,69 ✓		3

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RS Peters



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Holt Chemistry Ralph Thomas Myers,2004 *Elementary Modern Chemistry* Wilhelm Ostwald,Harry Wheeler Morse,1909 **Goyal's ICSE Chemistry Question Bank with Model Test Papers Class 10 for 2026 Examination** ,2025-07-02 *Goyal's I.C.S.E. Chemistry with Model Test Papers Class 10 for 2023 Examination* GBP Editorial,2022-08-10
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Modern Experimental Chemistry George W. Jr. Latimer,2012-12-02 Modern Experimental Chemistry provides techniques of qualitative analysis that reinforce experiments on ionic equilibriums This book includes the determination of water in hydrated salts identification of an organic compound after determining its molecular weight and nonaqueous titration of a salt of a weak acid The calculation of chemical stoichiometry calculation of thermodynamic properties by determining the change in equilibrium with temperature and chromium chemistry are also covered This compilation contains enough experiments for classes which have six hours of laboratory two 3 hour meetings per week to last two semesters This publication is intended for chemistry students as an introductory manual to chemistry laboratory **Goyal's ICSE**

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Inorganic Chemistry Joseph William Mellor,1917 **Holt Physical Science** William L. Ramsey,1997-11

Fundamentals of Chemistry: A Modern Introduction Frank Brescia,2012-12-02 Fundamentals of Chemistry Third Edition introduces the reader to the fundamentals of chemistry including the properties of gases atomic and molecular weights and the first and second laws of thermodynamics Chemical equations and chemical arithmetic are also discussed along with the structure of atoms chemical periodicity types of chemical bonds and condensed states of matter This book is comprised of 26 chapters and begins with a historical overview of chemistry and some terms which are part of the language of chemists Separation and purification are covered in the first chapter while the following chapters focus on atomic and

molecular weights stoichiometry the structure of atoms and types of chemical bonds The molecular orbital MO theory of bonding galvanic cells and chemical thermodynamics are considered next Separate chapters are devoted to MO theory of covalent and metallic bonding orbital hybridization intermolecular forces acids and bases ionic equilibrium calculations and polymers and biochemicals This monograph is intended for chemistry students [Molecular World](#) Catherine M. Jackson, 2023-07-04 A compelling and innovative account that reshapes our view of nineteenth century chemistry explaining a critical period in chemistry's quest to understand and manipulate organic nature According to existing histories theory drove chemistry's remarkable nineteenth century development In *Molecular World* Catherine M Jackson shows instead how novel experimental approaches combined with what she calls laboratory reasoning enabled chemists to bridge wet chemistry and abstract concepts and in so doing create the molecular world Jackson introduces a series of practice based breakthroughs that include chemistry's move into lampworked glassware the field's turn to synthesis and subsequent struggles to characterize and differentiate the products of synthesis and the gradual development of institutional chemical laboratories an advance accelerated by synthesis and the dangers it introduced Jackson's historical reassessment emerges from the investigation of alkaloids by German chemists Justus Liebig August Wilhelm Hofmann and Albert Ladenburg Stymied in his own research Liebig steered his student Hofmann into pioneering synthesis as a new investigative method Hofmann's practice based laboratory reasoning produced a major theoretical advance but he failed to make alkaloids That landmark fell to Ladenburg who turned to cutting edge theory only after his successful synthesis In telling the story of these scientists and their peers Jackson reveals organic synthesis as the ground chemists stood upon to forge a new relationship between experiment and theory with far reaching consequences for chemistry as a discipline **Study and Problem Solving Guide to Accompany Principles of Modern Chemistry, Oxtoby/Nachtrieb** Wade A. Freeman, 1986 **AP Chemistry Premium, 2025: Prep Book with 6 Practice Tests + Comprehensive Review + Online Practice** Barron's Educational Series, Neil D. Jespersen, Pamela Kerrigan, 2024-07-02 Be prepared for exam day with Barron's Trusted content from AP experts Barron's AP Chemistry Premium 2025 includes in depth content review and practice It's the only book you'll need to be prepared for exam day Written by Experienced Educators Learn from Barron's all content is written and reviewed by AP experts Build your understanding with comprehensive review tailored to the most recent exam Get a leg up with tips strategies and study advice for exam day it's like having a trusted tutor by your side Be Confident on Exam Day Sharpen your test taking skills with 6 full length practice tests 3 in the book and 3 more online plus 3 short diagnostic tests for assessing strengths and areas for improvement and detailed answer explanations for all questions Strengthen your knowledge with in depth review covering all units on the AP Chemistry exam Reinforce your learning with more than 300 practice questions throughout the book that cover all frequently tested topics Learn what to expect on test day with essential details about the exam format scoring calculator policy strategies for all question types and advice for developing a study plan Robust Online

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