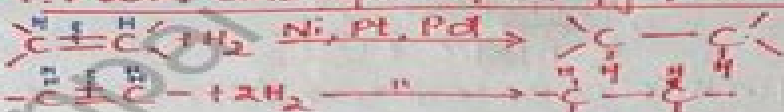


# HYDROCARBON

## # MOP of Alkane

### 1. From unsaturated hydrocarbon



DM  $\rightarrow$   $\pi$  bond ko H.H se tod do.

# exothermic

# 4 MCTS formed

# Syn addition

# Add<sup>n</sup> of conj. diene



#  $\text{RaNi} \rightarrow$  alloy of Ni & Al in NaOH  
Galvanom temp better than Ni (at 100°C)

# If  $\text{RaNi}$  used  $\rightarrow$  Sabatier Senderason's

#  $[\text{Rh}(\text{PPh}_3)_3\text{Cl}]$  used also

(Wilkinson's catalyst)

Alkene	Addition	Product
Cis	Syn	Erythro
Trans	Syn	Threo
Cis	Anti	Threo
Trans	Anti	Erythro

### 2. From alkyl halide



# X hatoo H lagao

#  $\text{ROR} : \text{R-Cl} < \text{R-Br} < \text{R-I}$   $\text{ROR} \propto \frac{1}{\text{BE}}$

### 3. Wurtz reaction



produce higher alkane with even no. of 'C'

#  $\text{R} - \text{X} \rightarrow$  coupling

#  $\text{R} - \text{X} \rightarrow$  disproportionation

# in 2 molecular wurtz faster than in 2 molecular but prod should be stable i.e A.A nahi banana hai.

#  $\text{ROR} \propto \frac{1}{\text{SH}}$   $\text{R-X} > 2^\circ \text{R-X} > 3^\circ \text{R-X}$

# for same R:  $\text{RF} < \text{RCl} < \text{RBr} < \text{RI}$

### 4. Decarboxylation reaction



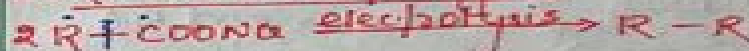
- COOH hatoo H lagao



- COONa hatoo H lagao

#  $\text{ROR} \propto$  stability of intermediate ( $\text{R}^\cdot$ )

### 5. Kolbe's elec. lysis



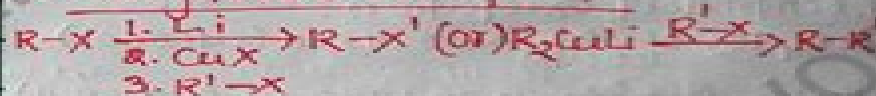
Cathode  $\rightarrow \text{H}_2$  Anode  $\rightarrow$  Alkane &  $\text{CO}_2$

### Addition Reaction

### 6. Frankland reaction



### 7. Coreyhouse reaction



# Attack of  $\text{R}^\cdot \rightarrow \text{RDS}$

#  $\text{ROR} \propto \frac{1}{\text{SH of R}' - \text{X}}$

# If same R:

$\text{R} - \text{F} < \text{R} - \text{Cl} < \text{R} - \text{Br} < \text{R} - \text{I}$

$\text{BE} \downarrow \rightarrow$  leaving tendency  $\uparrow$

# Mastery Study Guide Answers For Chemistry Hydrocarbons

**Brendan G. Carr**



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## **Table of Contents Mastery Study Guide Answers For Chemistry Hydrocarbons**

1. Understanding the eBook Mastery Study Guide Answers For Chemistry Hydrocarbons
  - The Rise of Digital Reading Mastery Study Guide Answers For Chemistry Hydrocarbons
  - Advantages of eBooks Over Traditional Books
2. Identifying Mastery Study Guide Answers For Chemistry Hydrocarbons
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Mastery Study Guide Answers For Chemistry Hydrocarbons
  - User-Friendly Interface
4. Exploring eBook Recommendations from Mastery Study Guide Answers For Chemistry Hydrocarbons
  - Personalized Recommendations
  - Mastery Study Guide Answers For Chemistry Hydrocarbons User Reviews and Ratings

- Mastery Study Guide Answers For Chemistry Hydrocarbons and Bestseller Lists
- 5. Accessing Mastery Study Guide Answers For Chemistry Hydrocarbons Free and Paid eBooks
  - Mastery Study Guide Answers For Chemistry Hydrocarbons Public Domain eBooks
  - Mastery Study Guide Answers For Chemistry Hydrocarbons eBook Subscription Services
  - Mastery Study Guide Answers For Chemistry Hydrocarbons Budget-Friendly Options
- 6. Navigating Mastery Study Guide Answers For Chemistry Hydrocarbons eBook Formats
  - ePub, PDF, MOBI, and More
  - Mastery Study Guide Answers For Chemistry Hydrocarbons Compatibility with Devices
  - Mastery Study Guide Answers For Chemistry Hydrocarbons Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Mastery Study Guide Answers For Chemistry Hydrocarbons
  - Highlighting and Note-Taking Mastery Study Guide Answers For Chemistry Hydrocarbons
  - Interactive Elements Mastery Study Guide Answers For Chemistry Hydrocarbons
- 8. Staying Engaged with Mastery Study Guide Answers For Chemistry Hydrocarbons
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Mastery Study Guide Answers For Chemistry Hydrocarbons
- 9. Balancing eBooks and Physical Books Mastery Study Guide Answers For Chemistry Hydrocarbons
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Mastery Study Guide Answers For Chemistry Hydrocarbons
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Mastery Study Guide Answers For Chemistry Hydrocarbons
  - Setting Reading Goals Mastery Study Guide Answers For Chemistry Hydrocarbons
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mastery Study Guide Answers For Chemistry Hydrocarbons
  - Fact-Checking eBook Content of Mastery Study Guide Answers For Chemistry Hydrocarbons
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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