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Steven Hester
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web indicate the type of reaction taking place $\text{NaOH} + \text{KNO}_3 \rightarrow \text{NaNO}_3 + \text{KOH}$ $\text{CH}_2\text{O} + \text{CO}_2 + \text{H}_2\text{O} \rightarrow 2\text{Fe} + 6\text{NaBr} + 2\text{FeBr}_6 + \text{Na}_3\text{MgOH} + \text{CaOH} + \text{MgSO}_4 + 2\text{NH}_4\text{OH} + \text{H}_2\text{O} + \text{NH}_4\text{Br} + \text{PbO}_2 + \text{Na}_2\text{CO}_3 + \text{Na}_2\text{O} + \text{CO}_2 + 4\text{H}_2\text{O} + 6\text{CO}_2 + 2\text{H}_2\text{O} + \text{HCl} + \text{NaOH} + \text{H}_2\text{O} + \text{NaCl}$

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taking place in the given diagram

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reaction a process in which one or more substances the reactants are converted to one or more different substances the
products substances are either chemical elements or compounds

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web to the left of the arrow we have the before situation this side represents the substances we have before the reaction
takes place they are called the reactants to the right of the arrow we have the after situation this side represents the
substances that we have after the reaction has taken place they are called the products

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chemical reaction single displacement 2 $1 \text{ CaCO}_3 \rightarrow 1 \text{ CaO} + 1 \text{ CO}_2$ type of chemical reaction decomposition

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2 $\text{Ca(OH)}_2 + \text{Al}_2\text{SO}_4 \rightarrow 3 \text{CaSO}_4 + \text{Al(OH)}_3$ type of reaction 3 $\text{Mg} + \text{Fe}_2\text{O}_3 \rightarrow \text{Fe} + \text{MgO}$ type of reaction 4 $\text{C}_2\text{H}_4 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ type of reaction 5
 $\text{PbSO}_4 + \text{PbSO}_3 + \text{O}_2$ type of reaction 6 NH_3 12

