

14.2 Math Skills

$$\text{Energy} = cm\Delta t$$

$$t_f = t_i - \frac{\text{energy}}{cm}$$

$$t_f = t_i + \frac{\text{energy}}{cm}$$

$$c = \frac{\text{energy}}{m\Delta t}$$

$$m = \frac{\text{energy}}{c\Delta t}$$

$$\Delta t = \text{---}$$

$$\Delta t = t_f - t_i$$

Specific Heat

☹ To receive any credit, you must show all your work, including givens, the correct equation, givens plugged in and answer with unit.

1. Brass is an alloy made from copper and zinc. A 0.59 kg brass candlestick has an initial temperature of 98.0°C. If 2.11×10^4 J of energy is removed from the candlestick to lower its temperature to 6.8°C, what is the specific heat of brass?
2. Mercury has one of the lowest specific heats. This fact added to its liquid state at most atmospheric temperatures makes it effective for use in thermometers. If 257 J of energy are added to 0.450 kg of mercury, the mercury's temperature will increase by 4.09 K. What is the specific heat of mercury?
3. Bismuth's specific heat is 121 J/kg • K, the lowest of any non-radioactive metal. What is the mass of a bismuth sample if 25 J raises its temperature 5.0 K?

Math Skills Specific Heat

AW Rasmussen

A red circular graphic with a gradient, appearing as a semi-circle or a stylized arrow pointing to the right, located on the right side of the slide.

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Table of Contents Math Skills Specific Heat

1. Understanding the eBook Math Skills Specific Heat
 - The Rise of Digital Reading Math Skills Specific Heat
 - Advantages of eBooks Over Traditional Books
2. Identifying Math Skills Specific Heat
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Math Skills Specific Heat
 - User-Friendly Interface
4. Exploring eBook Recommendations from Math Skills Specific Heat
 - Personalized Recommendations
 - Math Skills Specific Heat User Reviews and Ratings
 - Math Skills Specific Heat and Bestseller Lists
5. Accessing Math Skills Specific Heat Free and Paid eBooks
 - Math Skills Specific Heat Public Domain eBooks
 - Math Skills Specific Heat eBook Subscription Services
 - Math Skills Specific Heat Budget-Friendly Options
6. Navigating Math Skills Specific Heat eBook Formats

- ePub, PDF, MOBI, and More
- Math Skills Specific Heat Compatibility with Devices
- Math Skills Specific Heat Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Math Skills Specific Heat
 - Highlighting and Note-Taking Math Skills Specific Heat
 - Interactive Elements Math Skills Specific Heat
- 8. Staying Engaged with Math Skills Specific Heat
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Math Skills Specific Heat
- 9. Balancing eBooks and Physical Books Math Skills Specific Heat
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Math Skills Specific Heat
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Math Skills Specific Heat
 - Setting Reading Goals Math Skills Specific Heat
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Math Skills Specific Heat
 - Fact-Checking eBook Content of Math Skills Specific Heat
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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