#### Calorimetry: Heat of Neutralization

Prepared by Dongling Fei Manutee Community College

Purpose: Determine the calorimeter constant for a calorimeter. Determine the enthalpy of neutralization for the reaction of a strong acid and a strong base.

Background: When a system absorbs heat (q) its temperature changes by  $\Delta$  T. Experimental measurements demonstrate that the heat absorbed by a system and its corresponding temperature change are directly proportional. q=C x  $\Delta$  T. C is called the heat capacity, which is often defined as the quantity of heat required to change its temperature by  $1^{\circ}C$ 

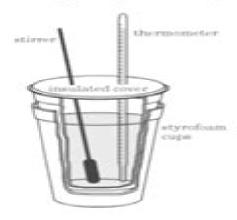
$$C = q/\Delta T = J/C$$

Heat capacity is an extensive property; it depends on the amount of matter being heated. The measure of intrinsic capacity of a substance to absorb heat is called **specific heat capacity** (C<sub>s</sub>), the amount of heat required to raise the temperature of 1 gram of the substance by 1 °C

$$C_s=q/(\Delta T *m)$$

The C<sub>s</sub> of H<sub>2</sub>O is 4.184J/g "C. For the case where all the heat is used to heat water, the amount of heat is calculated by the relation

The enthalpy change for a chemical reaction,  $\Delta H_{ma}$  is also called the **enthalpy of reaction** or **heat of reaction**, which is also an extensive property. For many aqueous reacions,  $\Delta H_{ma}$  can be measured simply using a coffee-cup calorimeter



# **Heat Of Neutralization Lab Manual**

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