Heat of Neutralization

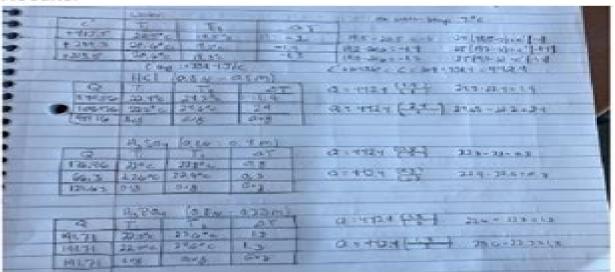
Introduction:

The experiment was used to determine the heat evolved in an aqueous acid-base reaction. For strong acids and bases reacting in sufficiently dilute solutions, the heat evolved is approximately 57 kJ mol-1 at 298 K. Increasing the concentration should increase the heat of the reaction, while using a weaker acid or base will reduce the heat of the reaction. A series of acid-base neutralizations were performed on aqueous sodium hydroxide and the heat of the reaction was determined calorimetrically.

Experimental:

To measure the heat capacity of the calorimeter system, 225 mL of distilled water was added to a clean, dry Dewar. The water was stirred and the temperature read every 20 seconds until it became constant. Then, after recording the temperature of it, 25 mL of ice water was added to the flask. The mixture was stirred and the temperature was read until constant readings were obtained. The Dewar flask was then dried and 200 mL of the 0.2 N NaOH solution was added to the calorimeter. The temperature was taken before 50 mL of the 0.8 N HCl solution was added. The mixture was stirred and the temperature was read until constant readings were obtained. This process was then repeated twice for both the sulfuric and phosphoric acid solutions.

Results:



Picture 1: Data tables and calculations for change in temperature and heat of neutralization

Heat Of Neutralization Lab Answer Key

Anthony S. Fauci

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Petroleum Engineer ,1954-07 ,1985 The Petroleum Engineer ,1954 Government Reports Announcements & **Petroleum Engineer for Management**, 1954 America's Textiles ,1977 **The Environment Index** Index ,1995 Transactions of the American Nuclear Society ,1984 Estimating a Heat of Neutralization M. L. Gillette, H. ,1986 Determination of the Heat of Neutralization Eugene Andrew Kaskiw, Arthur Joseph A. Neidig, J. N. Spencer, 1999-02-01 A Study of Methods for Determining Heats of Neutralization at Different Temperatures John Bernard Lavery, 1933 Laboratory Experiments for Chemistry Theodore E. Brown, John H. Nelson, Bruce E. Bursten, H. Eugene Dickson.1917 LeMay, Jr., Kenneth C. Kemp, Catherine Murphy, Patrick Woodward, 2019-11-12 This manual contains 43 finely tuned self contained experiments chosen to introduce basic lab techniques and to illustrate core chemical principles The Eleventh Edition has been revised to correlate more tightly with Brown LeMay Bursten's Chemistry The Central Science 11 e and now features a guide on how to keep a lab report notebook Safety and waste management are covered in greater detail and many pre lab and post lab questions have been updated The labs can also be customized through Catalyst Pearson's custom database program KEY TOPICS Basic Laboratory Techniques Identification of Substances by Physical Properties Separation of the Components of a Mixture Chemical Reactions Chemical Formulas Chemical Reactions of Copper and Percent Yield Chemicals in Everyday Life What Are They and How Do We Know Gravimetric Analysis of a Chloride Salt Gravimetric Determination of Phosphorus in Plant Food Paper Chromatography Separation of Cations and Dyes Molecular Geometries of Covalent Molecules Lewis Structures and the VSEPR model Atomic Spectra and Atomic Structure Behavior of Gases Molar Mass of a Vapor Determination of R The Gas Law Constant Activity Series Electrolysis the Faraday and Avogadro s Number Electrochemical Cells and Thermodynamics The Chemistry of Oxygen Basic and Acidic Oxides and the Periodic Table Colligative Properties Freezing Point Depression and Molar Mass Titration of Acids and Bases Reactions in Aqueous Solutions Metathesis Reactions and Net Ionic Equations Colorimetric Determination of an Equilibrium Constant in Aqueous Solution Chemical Equilibrium LeCh telier's Principle Hydrolysis of Salts and pH of Buffer Solutions Determination of the Dissociation Constant of a Weak Acid Titration Curves of Polyprotic Acids Determination of the Solubility Product Constant for a Sparingly Soluble Salt Heat of Neutralization Rates of Chemical Reactions I A Clock Reaction Rates of Chemical Reactions II Rate and Order of Decomposition Introduction to Qualitative Analysis Abbreviated Qualitative Analysis Scheme MARKET A hands on workbook CD useful for anyone studying general chemistry The Heat of Neutralization of Acids and Bases at 100° C. Alfred B. Babcock, 1909

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