



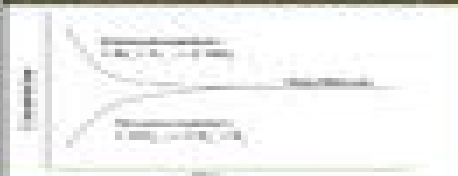
CHEMISTRY

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Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time (dynamic equilibrium). Equilibrium reactions are represented by the symbol $\rightleftharpoons$.

What is Equilibrium? Equilibrium is a state of balance between two or more opposing processes. In a chemical reaction, it is the state where the concentrations of the reactants and products remain constant over time. Equilibrium is a dynamic process, meaning that the reaction is still occurring, but the rates of the forward and reverse reactions are equal.	Equilibrium Constants K_c (Concentration) K_p (Pressure) K_a (Acid Dissociation) K_b (Base Dissociation)													
Equilibrium Constant, K_c For a general reaction: $aA + bB \rightleftharpoons cC + dD$ The equilibrium constant K_c is given by: $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	Equilibrium Constant, K_p <table><tr><th>Reaction</th><th>Equilibrium Constant</th></tr><tr><td>$N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$</td><td>$K_p = \frac{P_{NH_3}^2}{P_{N_2} P_{H_2}^3}$</td></tr><tr><td>$2SO_2(g) + O_2(g) \rightleftharpoons 2SO_3(g)$</td><td>$K_p = \frac{P_{SO_3}^2}{P_{SO_2}^2 P_{O_2}}$</td></tr><tr><td>$2H_2O(g) \rightleftharpoons 2H_2(g) + O_2(g)$</td><td>$K_p = \frac{P_{H_2}^2 P_{O_2}}{P_{H_2O}^2}$</td></tr></table>		Reaction	Equilibrium Constant	$N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$	$K_p = \frac{P_{NH_3}^2}{P_{N_2} P_{H_2}^3}$	$2SO_2(g) + O_2(g) \rightleftharpoons 2SO_3(g)$	$K_p = \frac{P_{SO_3}^2}{P_{SO_2}^2 P_{O_2}}$	$2H_2O(g) \rightleftharpoons 2H_2(g) + O_2(g)$	$K_p = \frac{P_{H_2}^2 P_{O_2}}{P_{H_2O}^2}$				
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Equilibrium Constant, K_a For a weak acid HA in water: $HA \rightleftharpoons H^+ + A^-$ The equilibrium constant K_a is given by: $K_a = \frac{[H^+][A^-]}{[HA]}$	Equilibrium Constant, K_b <table><tr><th>Reaction</th><th>Equilibrium Constant</th></tr><tr><td>$NH_3 + H_2O \rightleftharpoons NH_4^+ + OH^-$</td><td>$K_b = \frac{[NH_4^+][OH^-]}{[NH_3][H_2O]}$</td></tr><tr><td>$CH_3COO^- + H_2O \rightleftharpoons CH_3COOH + OH^-$</td><td>$K_b = \frac{[CH_3COOH][OH^-]}{[CH_3COO^-][H_2O]}$</td></tr><tr><td>$CO_3^{2-} + H_2O \rightleftharpoons HCO_3^- + OH^-$</td><td>$K_b = \frac{[HCO_3^-][OH^-]}{[CO_3^{2-}][H_2O]}$</td></tr></table>		Reaction	Equilibrium Constant	$NH_3 + H_2O \rightleftharpoons NH_4^+ + OH^-$	$K_b = \frac{[NH_4^+][OH^-]}{[NH_3][H_2O]}$	$CH_3COO^- + H_2O \rightleftharpoons CH_3COOH + OH^-$	$K_b = \frac{[CH_3COOH][OH^-]}{[CH_3COO^-][H_2O]}$	$CO_3^{2-} + H_2O \rightleftharpoons HCO_3^- + OH^-$	$K_b = \frac{[HCO_3^-][OH^-]}{[CO_3^{2-}][H_2O]}$				
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Le Chatelier's Principle: The effect of changing conditions on equilibrium <table><tr><th>Condition</th><th>Effect on Equilibrium</th><th>Example</th></tr><tr><td>Concentration</td><td> - Increasing concentration of a reactant shifts the equilibrium to the right. - Increasing concentration of a product shifts the equilibrium to the left. </td><td>$2SO_2(g) + O_2(g) \rightleftharpoons 2SO_3(g)$ Adding more SO_2 shifts the equilibrium to the right, producing more SO_3.</td></tr><tr><td>Pressure</td><td> - Increasing pressure shifts the equilibrium to the side with fewer moles of gas. - Decreasing pressure shifts the equilibrium to the side with more moles of gas. </td><td>$N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$ Increasing pressure shifts the equilibrium to the right, producing more NH_3.</td></tr><tr><td>Temperature</td><td> - Increasing temperature shifts the equilibrium to the endothermic side. - Decreasing temperature shifts the equilibrium to the exothermic side. </td><td>$2SO_2(g) + O_2(g) \rightleftharpoons 2SO_3(g)$ This reaction is exothermic. Increasing temperature shifts the equilibrium to the left, producing less SO_3.</td></tr></table>			Condition	Effect on Equilibrium	Example	Concentration	- Increasing concentration of a reactant shifts the equilibrium to the right. - Increasing concentration of a product shifts the equilibrium to the left.	$2SO_2(g) + O_2(g) \rightleftharpoons 2SO_3(g)$ Adding more SO_2 shifts the equilibrium to the right, producing more SO_3 .	Pressure	- Increasing pressure shifts the equilibrium to the side with fewer moles of gas. - Decreasing pressure shifts the equilibrium to the side with more moles of gas.	$N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$ Increasing pressure shifts the equilibrium to the right, producing more NH_3 .	Temperature	- Increasing temperature shifts the equilibrium to the endothermic side. - Decreasing temperature shifts the equilibrium to the exothermic side.	$2SO_2(g) + O_2(g) \rightleftharpoons 2SO_3(g)$ This reaction is exothermic. Increasing temperature shifts the equilibrium to the left, producing less SO_3 .
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web the `etc passwd` file is a colon separated file that contains the following information user name encrypted password user id number uid user s group id number gid full name of the user gecol user home directory login shell the following is an example of an `etc passwd` file

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