

CH 4. RADICALS

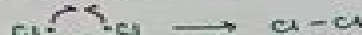
1. INITIATION - FORM RADICALS



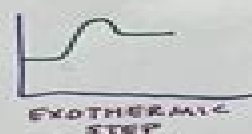
2. PROPAGATION - RADICAL REACTANT → RADICAL PRODUCT



3. TERMINATION - 2 RADICALS → NO RADICALS



$$\Delta H_{\text{rxn}} = \sum (\text{bonds broken}) - \sum (\text{bonds formed}) - \text{form bond}$$



F_2 - most reactive
 Cl_2 - reactive, not selective
 Br_2 - selective, picky
 I_2 - not reactive, endo.



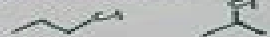
RESONANCE + SUBST. ↑ RAD. STABILITY
 $\text{Allyl} \cdot > \text{R}-\text{C}(\text{H})-\text{R} > \text{R}-\text{C}(\text{H})_2-\text{R} > \text{R}-\text{C}(\text{H})_3$
 $3^\circ > 2^\circ > 1^\circ$

CH 5. STEREOCHEMISTRY

CHIRAL - Not symmetrical



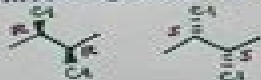
CONST. ISOMER - connect. diff.



ACHIRAL - SYMMETRY



ENANTIOMER - Invert chiral centers



DIASTEROMER - Invert some, not all, chiral centers



MESO - ACHIRAL W/ CHIRAL CENTERS



CONFORMATIONAL ISOMER

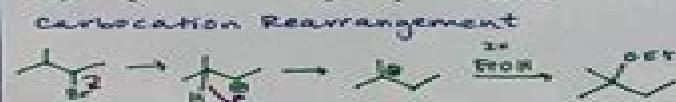


CH 6. NUCLEOPHILIC SUBSTITUTION



Nuc. Lone pairs (s) charge
 Electroph. (s) or s+ charge

SN1 → Racemic

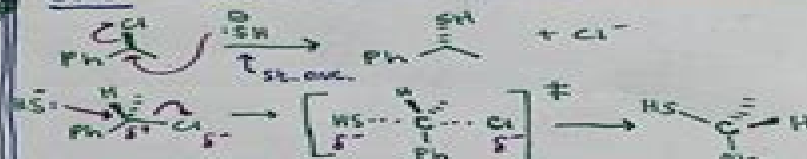


H or methyl shift, resonance
 * Weak nucleophile, Rate = k[subst.]

LEAVING GROUP: $\text{I} > \text{Br} > \text{Cl} \gg \text{F or OH}$



SN2 → Invert chiral center



* Strong Nucleophile
 1° or methyl
 rate = k[substrate][nucleophile]

Guide Quick Notes Reactions Iitjee Organic Chemistry

Dr Giriraj Vishwakarma



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