#### CHAPTER 7 REVIEW

# Chemical Formulas and Chemical Compounds

#### SECTION 7-1

SHORT ANSWER Answer the following questions in the space provided.

- In a Stock name such as iron(III) sulfate, the roman numeral tells us \_\_\_\_\_\_
  - (a) how many atoms of Fe are in one formula unit
  - (b) how many sulfate ions can be attached to the iron atom.
  - (c) the charge on each Fe ion
  - (d) the total positive charge of the formula unit
- The result of changing a subscript in a correctly written chemical formula is to \_\_\_\_\_\_
  - (a) change the number of moles represented by the formula
  - (b) change the charges on the other ions in the compound
  - (c) change the formula so that it no longer represents the compound it previously represented
  - (d) have no effect on the formula
- The explosive TNT has the molecular formula C<sub>2</sub>H<sub>3</sub>(NO<sub>2</sub>)<sub>3</sub>.
  - 4 elements a, How many elements make up this compound?
  - 6 oxygen atoms
    b. How many oxygen atoms are present in one molecule of C<sub>2</sub>H<sub>3</sub>(NO<sub>2</sub>)<sub>3</sub>?
  - 21 atoms c. How many atoms in total are present in one molecule of C<sub>7</sub>H<sub>5</sub>(NO<sub>2</sub>)<sub>3</sub>?
  - 4.2 × 10<sup>24</sup> atoms d. How many atoms are present in a sample of 2 × 10<sup>23</sup> molecules of C<sub>2</sub>H<sub>3</sub>(NO<sub>2</sub>)<sub>3</sub>?
- 4. How many atoms are present in each of these formula units?
  - 11 atoms a. Ca(HCO<sub>3</sub>)<sub>2</sub>
    - 45 atoms b. C<sub>12</sub>H<sub>22</sub>O<sub>11</sub>
    - 10 atoms c. Fe(ClO<sub>2</sub>)<sub>3</sub>
  - 9 atoms d. Fe(ClO<sub>3</sub>)<sub>2</sub>
- 5. N<sub>2</sub>O<sub>5</sub> a. What is the formula for the compound dinitrogen pentoxide?
  - carbon(IV) sulfide b. What is the Stock name for the covalent compound CS2?
    - H<sub>2</sub>SO<sub>3</sub> e. What is the formula for sulfurous acid?
    - phosphoric acid d. What is the name for the acid H<sub>3</sub>PO<sub>4</sub>?

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# **Guided And Study Chemical Names And Formulas**

**RS Peters** 

### **Guided And Study Chemical Names And Formulas:**

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