

General Chemistry II Jasperse  
Chemical equilibria. Extra Practice Problems

General Types/Groups of problems:

|   |    |  |     |
|---|----|--|-----|
| Equilibrium Conceptual  | p1 | Using Iee: Generic, Then Real But Simple Numbers                       | p8  |
| Writing the Equilibrium Constant  | p3 | Solving for K given Initial and at Least one Equilibrium Concentration | p9  |
| Manipulations of K: Reversing or Multiplying  | p5 | Solving for Equilibrium Concentrations Using Iee, Given K              | p10 |
| Solving for K Given All Equilibrium Concentrations                                    | p6 | LeChatelier's Principle, Q, and Responses to Disrupted Equilibria      | p12 |
| Solving for an Equilibrium Concentration Given K and Other Equilibrium Concentrations | p7 | Answers  | p15 |

**Big-Picture Introductory Conceptual Questions**

- Which of the following is true for a chemical reaction at equilibrium?
  - only the forward reaction stops
  - only the reverse reaction stops
  - both the forward and reverse reactions stop
  - the rate constants for the forward and reverse reactions are equal
  - the rates of the forward and reverse reactions are equal
- Which of the following is true regarding the **concentration of products**, for a chemical reaction **that is already at equilibrium**, assuming no disruptions to the equilibrium?
  - The concentrations of products will not change because there are no more reactants.
  - The concentrations of products will not change because the limiting reagent is gone.
  - The concentrations of products will not change because the forward and reverse rates are equal.
  - The concentrations of products will change continually because of reversibility.
- Which of the following are equal for a chemical system at equilibrium? If all are equal, answer E.
  - the **concentrations** of reactant and products are equal
  - the **rate constants** for the forward and reverse reactions are equal
  - the **time** that a particular atom or molecule spends as a reactant and product are equal
  - the **rate** of the forward and reverse reaction
  - all** of the above are equal
- A chemical equilibrium may be established by starting a reaction with \_\_\_\_\_.
  - reactants only.
  - products only.
  - equal quantities of reactants and products.
  - any quantities of reactants and products.
  - all the above
- An equilibrium that strongly favors products has \_\_\_\_\_.
  - a value of  $K \ll 1$ .
  - a value of  $K \gg 1$ .
  - a value of  $Q \gg 1$ .
  - a value of  $Q \ll 1$ .
  - $K = Q$ .

# Guided Practice Problems Chemistry Ch 16

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**Survival Guide to Organic Chemistry** Patrick E. McMahon, Bohdan B. Khomtchouk, Claes Wahlestedt, 2016-12-19 Reviews key general chemistry concepts and techniques adapted for application to important organic principles Provides practical guidance to help students make the notoriously well known and arduous transition from general chemistry to organic chemistry Explains organic concepts and reaction mechanisms generally expanding the focus on how to understand each step from a more intuitive viewpoint Covers concepts that need further explanation as well as those that summarize and emphasize key ideas or skills necessary in this field An added bonus is help with organizing principles to make sense of a wide range of similar reactions and mechanisms Implements a user friendly process to achieve the end result of problem solving Covers organic chemistry I and II concepts at the level and depth of a standard ACS organic chemistry curriculum features practice problems and solutions to help master the material including an extensive and comprehensive bank of practice exams with solutions

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## **Table of Contents Guided Practice Problems Chemistry Ch 16**

1. Understanding the eBook Guided Practice Problems Chemistry Ch 16
  - The Rise of Digital Reading Guided Practice Problems Chemistry Ch 16
  - Advantages of eBooks Over Traditional Books
2. Identifying Guided Practice Problems Chemistry Ch 16
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Guided Practice Problems Chemistry Ch 16
  - User-Friendly Interface
4. Exploring eBook Recommendations from Guided Practice Problems Chemistry Ch 16
  - Personalized Recommendations
  - Guided Practice Problems Chemistry Ch 16 User Reviews and Ratings
  - Guided Practice Problems Chemistry Ch 16 and Bestseller Lists

5. Accessing Guided Practice Problems Chemistry Ch 16 Free and Paid eBooks
  - Guided Practice Problems Chemistry Ch 16 Public Domain eBooks
  - Guided Practice Problems Chemistry Ch 16 eBook Subscription Services
  - Guided Practice Problems Chemistry Ch 16 Budget-Friendly Options
6. Navigating Guided Practice Problems Chemistry Ch 16 eBook Formats
  - ePub, PDF, MOBI, and More
  - Guided Practice Problems Chemistry Ch 16 Compatibility with Devices
  - Guided Practice Problems Chemistry Ch 16 Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Guided Practice Problems Chemistry Ch 16
  - Highlighting and Note-Taking Guided Practice Problems Chemistry Ch 16
  - Interactive Elements Guided Practice Problems Chemistry Ch 16
8. Staying Engaged with Guided Practice Problems Chemistry Ch 16
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Guided Practice Problems Chemistry Ch 16
9. Balancing eBooks and Physical Books Guided Practice Problems Chemistry Ch 16
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Guided Practice Problems Chemistry Ch 16
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Guided Practice Problems Chemistry Ch 16
  - Setting Reading Goals Guided Practice Problems Chemistry Ch 16
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Guided Practice Problems Chemistry Ch 16
  - Fact-Checking eBook Content of Guided Practice Problems Chemistry Ch 16
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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