

Isotopes Practice Set

- What does atomic number tell you about an element?
- What does the number used to isotopes signify?
- How can you tell isotopes apart?

For each of the following isotopes, write the number of protons, neutrons, and electrons.

	Chlorine-35	Chlorine-37
# of protons	17	17
# of neutrons	18	20
# of electrons	17	17

	Nitrogen-14	Nitrogen-15
# of protons	7	7
# of neutrons	7	8
# of electrons	7	7

	Carbon-12	Carbon-13
# of protons	6	6
# of neutrons	6	7
# of electrons	6	6

	Sulfur-32	Sulfur-34
# of protons	16	16
# of electrons	16	16
# of neutrons	16	18

Fill in the isotopes names including chemical symbol, atomic number, and mass number from the values that your partner gave and the information provided.

	Isotope: 15	Isotope: 23
# of protons	15	15
# of neutrons	8	8
# of electrons	15	15

	Isotope: 20	Isotope: 27
# of protons	20	20
# of neutrons	10	10
# of electrons	20	20

	Isotope: 14	Isotope: 16
# of protons	14	14
# of neutrons	7	7
# of electrons	14	14

	Isotope: 12	Isotope: 14
# of protons	12	12
# of neutrons	6	6
# of electrons	12	12

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