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# Gmc Jimmy 96 Repair Manual

**Katrin Zwirgmaier**



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web 1 define each remember oil rig oxidation is loss of e reduction is gain of e a oxidation b reduction c oxidizing agent d reducing agent 2 write half reactions for each of the following atoms or ions label each as oxidation or reduction use the activity series chart a al b  $\text{Ba}^{2+}$  c  $\text{Br}_2$  d ca e  $\text{Ga}^{3+}$  f h 2 g h

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web you should try to answer the questions without referring to your textbook if you get stuck try asking another group for help 1 identify the oxidation state of each element in the following a na 2 so 4 b mno 4 c kclo 3 d ch 3 oh 2 identify the oxidation state of nitrogen in the following a no 3 b no 2 c no 2

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web write the oxidation and reduction reactions for each redox reaction the first one is done for you 23  $\text{Fe}^{2+} + \text{Co} \rightarrow \text{Co}^{2+} + \text{Fe}$   
oxidation  $\text{Co} \rightarrow \text{Co}^{2+} + 2\text{e}^{-}$  reduction  $\text{Fe}^{2+} + 2\text{e}^{-} \rightarrow \text{Fe}$  24 3  $\text{Ag} + \text{Ni}^{3+} \rightarrow \text{Ni} + 3\text{Ag}^{+}$  oxidation reduction 25  $\text{Cu}^{2+} + \text{Pb} \rightarrow \text{Pb}^{2+} + \text{Cu}$  oxidation reduction 26  $\text{O}_2 + 2\text{H}_2 \rightarrow 2\text{H}_2\text{O}$

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web in each of the following redox reactions indicate the elements that undergo changes in oxidation number and the magnitude of change then identify the element oxidized and reduced as well as the oxidizing agent and reducing agent a  $\text{I}_2\text{O}_5 + 5\text{CO} \rightarrow \text{I}_2 + 5\text{CO}_2$  b  $\text{I}_2 + 2\text{O}_5 \rightarrow \text{I}_2\text{O}_5 + 2\text{CO}_2$

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web chemical reactions in which there is a transfer of electrons from one substance to another are known as oxidation reduction reactions or redox reactions in this unit you will examine the oxidation reduction process and use the oxidation

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web jun 28 2010 answers i determine what is oxidized and what in each reaction identify the oxidizing agent and the reducing agent also  $2\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$   $\text{Sr}^0$  to  $\text{Sr}^{+2}$  oxidized reducing agent  $\text{O}^0$  to  $\text{O}^{-2}$  reduced ox ag  $2\text{Li} + \text{S} \rightarrow \text{Li}_2\text{S}$   $\text{Li}^0$  to  $\text{Li}^{+1}$  oxidized red ag  $\text{S}^0$  to  $\text{S}^{-2}$  reduced ox ag  $3\text{Cs} + \text{Br}_2$

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