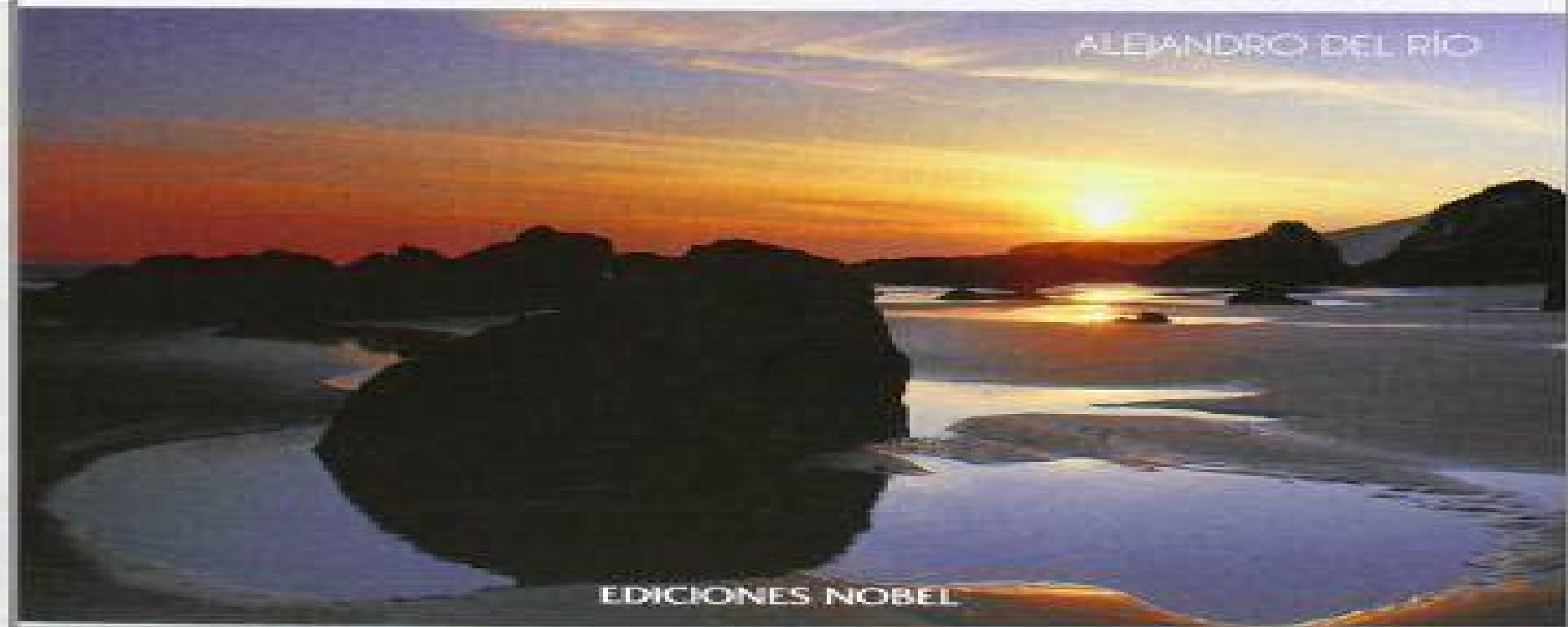




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$$4\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Cu} + 2\text{Ag} + 2\text{CuNO}_3 + 2\text{AgNO}_3$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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$$8.3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} = 8.3 \text{ mole H}_2\text{O}$$

8.3 mole H₂S x 18g H₂O = 149.4g H₂O is your theoretical yield so percent yield = $\frac{137.1\text{g}}{149.4\text{g}} \times 100 = 91.77\%$

problems 6 & 8 are similar to problem 5

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$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$

if the reaction of 21.3 grams of O₂ produces a 47.3% yield how many grams of CO would be produced

step one figure out how many moles of the limiting reagent you have

$$\frac{21.3 \text{ g O}_2}{32 \text{ g/mol}} = 0.666 \text{ moles O}_2$$

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web percent yield is a measure of the efficiency of the experimental design

$$\text{yield} = \frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$$

in this reaction a yield of 80% is anticipated

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web a write a balanced equation for the reaction

$$4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$$

b determine the theoretical yield of Al₂O₃

$$160.0 \text{ g Al} \times \frac{1 \text{ mol Al}}{27.0 \text{ g/mol}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} \times 101.96 \text{ g/mol} = 150.0 \text{ g Al}_2\text{O}_3$$

c determine the percent yield

$$\frac{126.0 \text{ g}}{150.0 \text{ g}} \times 100 = 84.0\%$$

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