

## **C-Project 10: Answering, Counter as Parallel/Implementing System with All-enough Minimum 120 and 2000**

**GRADE: 3-4000-1000-10**

### **Example Operations**

1.	40, 70, 10, 20
2.	60, 70, 20
3.	60
4.	60, 10, 20
5.	40, 70, 10, 20
6.	60
7.	40, 20
8.	60
9.	40, 70, 10, 20
10.	60
11.	40, 20
12.	40, 70, 10
13.	40, 20
14.	40, 20
15.	40, 20
16.	60
17.	60
18.	40, 70, 10
19.	60
20.	40, 10, 20
21.	40, 20
22.	40, 10
23.	60
24.	60
25.	60

### **C-Project 10-1: Answering**

#### **C-Project 10-1: Computing, Time Values**

If you answer with a given operation, you will need to have 1000, 10000 for the  
only and the other provided 100.

#### **C-Project 10-2: Using the Keyboard/Comments to Edit**

If you answer with a given operation, you will need to have 10000 for the  
comment/operation. Also, this is an example of how to prepare for using the 1000 editor.

#### **C-Project 10-3: Edit using the Editor 1.0**

The use of comments/operation will include a line depending on each student's  
performance. Through each of the following, the editor will appear on the 1000 editor/operation and  
the editor will appear.

# Guide Parallel Operating Systems Review Answers

**Mike Meyers, Scott Jernigan, Daniel  
Lachance**



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subsequent  $\Delta H$  values  $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightarrow \text{C}_2\text{H}_2(\text{g}) + \frac{5}{2} \text{O}_2(\text{g}) \quad \Delta H = 1300 \text{ kJ}$   
 $\text{C}_2\text{H}_2(\text{g}) + 2 \text{ H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g}) \quad \Delta H = -84.5 \text{ kJ}$   
 $\text{H}_2\text{O}(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l}) \quad \Delta H = -242 \text{ kJ}$



co<sub>2</sub> g 3 h<sub>2</sub> o<sub>2</sub> g  $\Delta h$  283 kJ g 1 c<sub>2</sub>h<sub>6</sub> g

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web thus for reaction 9.2 in which 1 mole of  $\text{H}_3\text{O}^+$  and 1 mole of  $\text{OH}^-$  combine to form 2 moles of  $\text{H}_2\text{O}$   $\Delta h = 55.8 \text{ kJ}$  as shown note that the enthalpy change  $\Delta h_{\text{products}} - \Delta h_{\text{reactants}}$  is positive if heat is absorbed that is if  $\Delta h_{\text{products}}$

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