

its second half equivalence point?

- A. $[\text{CO}_3^{2-}] = \text{maximum value}$
- B. $[\text{HCO}_3^-] = \text{maximum value}$
- C. $[\text{H}_2\text{CO}_3] = [\text{HCO}_3^-]$
- D. $[\text{HCO}_3^-] = [\text{CO}_3^{2-}]$
- E. $[\text{H}_2\text{CO}_3] = [\text{CO}_3^{2-}]$

Answer: D. $[\text{HCO}_3^-] = [\text{CO}_3^{2-}]$

4. The K_a for HCN is 6.2×10^{-10} , If there is a solution of 2M HCN, what concentration of NaCN is needed in order for the pH to be 9.2?

- A. 3M
- B. No NaCN needs to be added
- C. 2M
- D. 1M

Answer: C. 2M

5. NaOH is added to a 500ml of 2M acetic acid. If the $\text{p}K_a$ value of acetic acid is approximately 4.8, what volume of 2M NaOH must be added so that the pH of the solution is 4.8?

- A. 2L
- B. 500ml
- C. 250mL
- D. 1L

Answer: C. 250ml

6. Consider the titration curve, in which an acid is treated with aqueous 0.5M NaOH. What is the approximate $\text{p}K_a$ of the starting acid?

- A. 4.2
- B. 3.0
- C. 8.9
- D. 5.0
- E. 7.0

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